

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.94 g of iron and 35.48 g of sulfur. Find the empirical formula.

$$\text{40.94 g Fe} \times \frac{1 \text{ mol Fe}}{55.85 \text{ g Fe}} = 0.7332 \text{ mol Fe}$$

$$\text{35.48 g S} \times \frac{1 \text{ mol S}}{32.06 \text{ g S}} = 1.107 \text{ mol S}$$

FeS

2. A compound is found to contain 20.0 g of carbon, 26.7 g of oxygen, and 6.7 g of hydrogen. Find the empirical formula.

$$\text{20.0 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.66 \text{ mol C}$$

$$\text{26.7 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O}$$



$$\text{6.7 g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.63 \text{ mol H} \quad \text{C}_1\text{O}_1\text{H}_{6.63} \rightarrow \text{C}_1\text{O}_1\text{H}_7$$

3. A white solid is found to contain 40.94 g of iron, 35.48 g of sulfur, 20.0 g of carbon, 26.7 g of oxygen, and 6.7 g of hydrogen. Find the empirical formula.

$$\text{40.94 g Fe} \times \frac{1 \text{ mol Fe}}{55.85 \text{ g Fe}} = 0.7332 \text{ mol Fe}$$



$$\text{35.48 g S} \times \frac{1 \text{ mol S}}{32.06 \text{ g S}} = 1.107 \text{ mol S} \quad \text{Fe-C}_1\text{O}_1\text{H}_7\text{S}_1$$

$$\text{20.0 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.66 \text{ mol C} \quad \text{Fe-C}_1\text{O}_1\text{H}_7\text{S}_1\text{C}_{1.66}$$

$$\text{26.7 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O} \quad \text{Fe-C}_1\text{O}_1\text{H}_7\text{S}_1\text{C}_{1.66}\text{O}_{1.67}$$

4. A 100.00 g compound is analyzed and found to contain 20.0 g carbon, 26.7 g oxygen, and 6.7 g hydrogen. Find the empirical formula.

$$\text{20.0 g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.66 \text{ mol C}$$



$$\text{26.7 g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O}$$

$$\text{6.7 g H} \times \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 6.63 \text{ mol H} \quad \text{C}_1\text{O}_1\text{H}_7\text{S}_1\text{C}_{1.66}\text{O}_{1.67}$$

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