Empirical and Molecular Formula Worksheet ANSWER KEY

Write the empirical formula for the following compounds.

- D C.H. CH
- 6) C₂H₁₁ C₂H₂
- 7) WO. WO.
- 8) C2H4O2 CH.O
- 9) X2, Y ... X, Y
- 6) A compound with an empirical formula of C₂OH₄ and a molar mass of 88 grams per mole. What is the molecular formula of this compound? C₄O₂H₄.
- 7) A compound with an empirical formula of C_aH_aO and a molar mass of 136 grams per mole. What is the molecular formula of this compound? C_aH_aO_a
- 8) A compound with an empirical formula of CFBrO and a molar mass of 253.8 grams per mole. What is the molecular formula of this compound? C₂F₂Br₂O₃.
- 9) A compound with an empirical formula of C₂H₆N and a molar mass of 46 grams per mole. What is the molecular formula of this compound? C₂H₆N
- 10) A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula C₂H₂NO. If its molar mass is 116.1 g/mol, what is the molecular formula of the compound? C₂H₂N₂O₃.
- 12. A certain blue solid contains 36.84% N and 63.16% O. What is the empirical formula of this compound? The ratios are N_{1.00}O_{1.50}. Since 1.50 is not close to a whole number, we multiply both subscripts by 2. The empirical formula is thus N₂O₃. (The name is dinitrogen trioxide.)
- A sample of indium chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the indium compound? InCl,
- 14. An unknown compound was found to have a percent composition as follows: 47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula? If the true molar mass of the compound is 166.22 g/mol, what is its molecular formula? K₂C₂O₄
- Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of subbing alcohol? C₃H₄O

Empirical Formula Answer Key

Arihant Experts

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