

# Exercise 1: Kinetics

1.1. A reaction with reactants and products. Consider the reaction with the following forward and reverse rate constants:

1. No forward reaction (rate constant  $k_f = 0$ ) and reverse reaction (rate constant  $k_r = 1$ ).

2. Both  $k_f = 1$  and  $k_r = 1$  (equilibrium).

3. Both  $k_f = 1$  and  $k_r = 1$  (equilibrium).

1.2. For each reaction, write the corresponding rate equations for the forward and reverse reactions and the overall rate of reaction.

1. Forward reaction rate  $k_f = 0$  and reverse reaction rate  $k_r = 1$ .

2. Forward reaction rate  $k_f = 1$  and reverse reaction rate  $k_r = 1$ .

3. Forward reaction rate  $k_f = 1$  and reverse reaction rate  $k_r = 1$ .

1.3. For each reaction, write the corresponding rate equations for the forward and reverse reactions and the overall rate of reaction.

1. Forward reaction rate  $k_f = 0$  and reverse reaction rate  $k_r = 1$ .

2. Forward reaction rate  $k_f = 1$  and reverse reaction rate  $k_r = 1$ .

3. Forward reaction rate  $k_f = 1$  and reverse reaction rate  $k_r = 1$ .

1.4. Write the rate equations for the reaction  $2\text{H}_2\text{O} \rightarrow \text{H}_2 + \text{O}_2$ .

1. Forward reaction rate  $k_f = 1$  and reverse reaction rate  $k_r = 1$ .

2. Forward reaction rate  $k_f = 1$  and reverse reaction rate  $k_r = 1$ .

3. Forward reaction rate  $k_f = 1$  and reverse reaction rate  $k_r = 1$ .

1.5. Write the rate equations for the reaction  $2\text{H}_2\text{O} \rightarrow \text{H}_2 + \text{O}_2$ .

1. Forward reaction rate  $k_f = 1$  and reverse reaction rate  $k_r = 1$ .

2. Forward reaction rate  $k_f = 1$  and reverse reaction rate  $k_r = 1$ .

# Lab Manual Answers Spanish Imagina Leccion 6

**Léo Pomar**



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