

Types of Reactions Lab

SAFETY

- When heating a test tube, aim the mouth of a test tube AWAY from yourself and others.
- Do NOT look directly at burning magnesium. The intense light may damage your eyes.
- Hydrochloric acid can cause burns to the skin and damage the eyes.

PROCEDURE

Reaction A

1. Add about 1 mL (1 mL = width of your pinky finger) of calcium chloride solution, CaCl_2 , to a clean test tube. Next, add about 1 mL of sodium carbonate solution, Na_2CO_3 , to the same test tube. Record your observations.

Reaction B

2. Stand a clean test tube in the test tube rack and add 1-2 mL of hydrochloric acid. CAUTION: HCl can cause burns! Carefully drop a few pieces of magnesium into test tube.
3. While the reaction is still occurring, use a test tube holder to hold an empty test tube over the top of the first test tube for ~30 seconds. Meanwhile, another group member should light the burner and use it to ignite a wood splint. At the end of 30 seconds, hold the inverted test tube (keeping it inverted) while the other group member places the burning wood splint into the mouth of the inverted test tube to test for the presence of gas. Record your observations, including the results of the hydrogen test. KEEP THE BURNER LIT FOR REACTION C.

Reaction C

4. Place two spatula tips full of sodium hydrogen carbonate, NaHCO_3 , in a clean test tube. Note the appearance of the sample in your observations.
5. Using a test tube holder, heat the NaHCO_3 strongly in the burner flame for 2-3 minutes. AIM THE MOUTH OF THE TEST TUBE AWAY FROM YOURSELF AND OTHERS. After heating, another group member should ignite a wood splint and quickly place the burning splint into the test tube to test for the presence and type of gas.
6. Record your observations, including the results of the gas test. KEEP THE BURNER LIT FOR REACTION D.

Reaction D

7. Place a watch glass next to the burner. Examine a piece of magnesium ribbon and record your observations. Using crucible tongs, hold the ribbon in the burner flame until the magnesium starts to burn. DO NOT LOOK DIRECTLY AT THE FLAME. Hold the burning magnesium directly over the watch glass. When the ribbon stops burning, place the remains on the watch glass. Examine this product thoroughly and record your observations. Turn off the burner.
8. Clean your test tubes and lab station. Wash liquids down the drain. Wrap solids in a paper towel before discarding them in the trash can.

OBSERVATIONS

Record the appearance of the reactants before and the products after each reaction takes place. Make a chart in your lab notebook.

ANALYSIS

Write the chemical equations for the reactions that occurred in each step. Balance the reaction and state what type each is. Each equation should be recorded in formulas (symbols) and words.

CONCLUSIONS

Answer the following questions based on the results of your lab:

1. What signs did you observe that indicated a chemical reaction was taking place? Be specific.
2. The test for gas produces different results for H_2 , O_2 and CO_2 gases. What result does each produce and why? (Hint: think back to the rocket lab).

HONORS: A combustion reaction was taking place each time you used the Bunsen burner to burn methane gas (CH_4). Write a balanced chemical equation for the combustion of methane

Lab Manual For High School Chemistry

CH Cherryholmes



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