

Experiment 10

Oxidation-Reduction (Redox) Reactions

Pre-Lab Assignment

Before coming to lab:

- Read the lab thoroughly.
- Answer the pre-lab questions that appear at the end of this lab exercise.

Purpose

A set of oxidation-reduction (redox) reactions will be performed to determine the relative strengths as reducing and oxidizing agents of metals and other ions. An activation series will be devised and used to predict whether a reaction will occur spontaneously or not.

Background

Oxidation-reduction (redox) reactions are a classification of chemical changes that involve the transfer of electrons. An example of a redox reaction is shown in Eqn. 1, when magnesium metal reacts with chlorine gas.



At first glance, the electron exchange is not apparent as electrons are not included in the overall balanced equation. However, electrons are shown when the reaction is split into halves: an oxidized half-reaction and a reduced half-reaction, as in Eqn. 2 and Eqn. 3 respectively.



Notice that in the half-reactions, the electrons are required to balance the charges. Since Mg is becoming more positive, one magnesium atom loses two electrons to become a cation and is said to be **oxidized**. Chlorine is becoming more negative, so one molecule of $\text{Cl}_2(\text{g})$ gains two electrons to become two chloride anions and is said to be **reduced**. Both oxidation and reduction happen simultaneously in a single redox reaction. Overall, Mg is transferring two electrons to chlorine (Fig. 1).

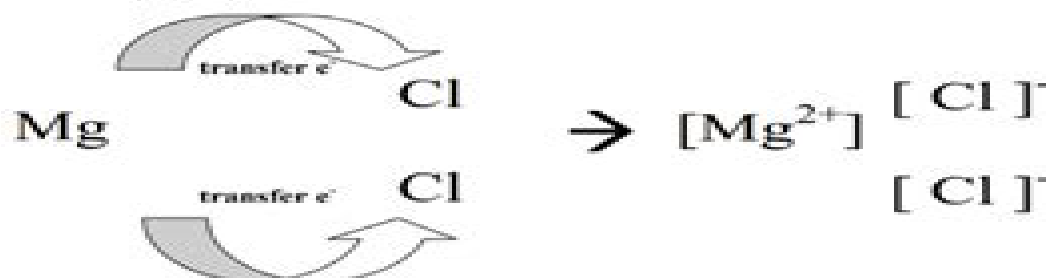


Fig. 1: Redox reaction between Mg(s) and Cl₂(g) to produce MgCl₂(aq)

Lab Manual Answers Redox Oxidation Reduction Reaction

BM King



Lab Manual Answers Redox Oxidation Reduction Reaction:

The book delves into Lab Manual Answers Redox Oxidation Reduction Reaction. Lab Manual Answers Redox Oxidation Reduction Reaction is a vital topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Lab Manual Answers Redox Oxidation Reduction Reaction, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Lab Manual Answers Redox Oxidation Reduction Reaction
 - Chapter 2: Essential Elements of Lab Manual Answers Redox Oxidation Reduction Reaction
 - Chapter 3: Lab Manual Answers Redox Oxidation Reduction Reaction in Everyday Life
 - Chapter 4: Lab Manual Answers Redox Oxidation Reduction Reaction in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, this book will provide an overview of Lab Manual Answers Redox Oxidation Reduction Reaction. The first chapter will explore what Lab Manual Answers Redox Oxidation Reduction Reaction is, why Lab Manual Answers Redox Oxidation Reduction Reaction is vital, and how to effectively learn about Lab Manual Answers Redox Oxidation Reduction Reaction.
 3. In chapter 2, the author will delve into the foundational concepts of Lab Manual Answers Redox Oxidation Reduction Reaction. The second chapter will elucidate the essential principles that need to be understood to grasp Lab Manual Answers Redox Oxidation Reduction Reaction in its entirety.
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