

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization energy) increases.

Lab Periodic Trends Key

Rachel S Tattersall



Lab Periodic Trends Key:

Unveiling the Energy of Verbal Artistry: An Mental Sojourn through **Lab Periodic Trends Key**

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Table of Contents Lab Periodic Trends Key

1. Understanding the eBook Lab Periodic Trends Key
 - The Rise of Digital Reading Lab Periodic Trends Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Periodic Trends Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Periodic Trends Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Periodic Trends Key
 - Personalized Recommendations

- Lab Periodic Trends Key User Reviews and Ratings
- Lab Periodic Trends Key and Bestseller Lists
- 5. Accessing Lab Periodic Trends Key Free and Paid eBooks
 - Lab Periodic Trends Key Public Domain eBooks
 - Lab Periodic Trends Key eBook Subscription Services
 - Lab Periodic Trends Key Budget-Friendly Options
- 6. Navigating Lab Periodic Trends Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Periodic Trends Key Compatibility with Devices
 - Lab Periodic Trends Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Periodic Trends Key
 - Highlighting and Note-Taking Lab Periodic Trends Key
 - Interactive Elements Lab Periodic Trends Key
- 8. Staying Engaged with Lab Periodic Trends Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Periodic Trends Key
- 9. Balancing eBooks and Physical Books Lab Periodic Trends Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Periodic Trends Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lab Periodic Trends Key
 - Setting Reading Goals Lab Periodic Trends Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lab Periodic Trends Key
 - Fact-Checking eBook Content of Lab Periodic Trends Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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