

1. What volume of hydrogen gas at STP is produced from the reaction of 50.0g of Mg and 75.0 grams of HCl? How much of the excess reagent is left over (in grams)?



Do a standard mass to mass problem starting with each reactant

50.0 g Mg	$\frac{1 \text{ mol Mg}}{24.31 \text{ g Mg}}$	$\frac{1 \text{ mol H}_2}{1 \text{ mol Mg}}$	$\frac{22.4 \text{ L}}{1 \text{ mol H}_2}$	= 46.1 L H ₂
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75.0 g HCl	$\frac{1 \text{ mol HCl}}{36.46 \text{ g HCl}}$	$\frac{1 \text{ mole H}_2}{2 \text{ mol HCl}}$	$\frac{22.4 \text{ L}}{1 \text{ mol H}_2}$	= 23.0 L H ₂
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HCl is the limiting reactant!!

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J Dewey



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