

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

|                                             |                         |
|---------------------------------------------|-------------------------|
| _____ <b>lattice energy</b> _____           | <b>a.</b> ionic bond    |
| _____ <b>bond energy</b> _____              | <b>b.</b> covalent bond |
| _____ <b>enthalpy of vaporization</b> _____ | <b>c.</b> metallic bond |

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

|                                   |                 |                                      |               |
|-----------------------------------|-----------------|--------------------------------------|---------------|
| _____ <b>ionic</b> _____          | <b>a.</b> H—F   | _____ <b>nonpolar covalent</b> _____ | <b>d.</b> H—H |
| _____ <b>ionic</b> _____          | <b>b.</b> Na—Cl | _____ <b>polar covalent</b> _____    | <b>e.</b> H—C |
| _____ <b>polar covalent</b> _____ | <b>c.</b> H—O   | _____ <b>polar covalent</b> _____    | <b>f.</b> H—N |

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen  
atom and the unshared electron pair of a strongly electronegative atom such as O,  
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,  
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally  
intermolecular, while ionic and covalent bonds occur between ions or atoms  
respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.  
Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making  
water a liquid at room temperature. Hydrogen bonding exists between water  
molecules, but not between hydrogen sulfide molecules.

# Glencoe Covalent Bonding Chapter 8 Worksheet

**Thomas Griffiths**



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### **Table of Contents Glencoe Covalent Bonding Chapter 8 Worksheet**

1. Understanding the eBook Glencoe Covalent Bonding Chapter 8 Worksheet
  - The Rise of Digital Reading Glencoe Covalent Bonding Chapter 8 Worksheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Glencoe Covalent Bonding Chapter 8 Worksheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Glencoe Covalent Bonding Chapter 8 Worksheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Glencoe Covalent Bonding Chapter 8 Worksheet

- Personalized Recommendations
  - Glencoe Covalent Bonding Chapter 8 Worksheet User Reviews and Ratings
  - Glencoe Covalent Bonding Chapter 8 Worksheet and Bestseller Lists
5. Accessing Glencoe Covalent Bonding Chapter 8 Worksheet Free and Paid eBooks
    - Glencoe Covalent Bonding Chapter 8 Worksheet Public Domain eBooks
    - Glencoe Covalent Bonding Chapter 8 Worksheet eBook Subscription Services
    - Glencoe Covalent Bonding Chapter 8 Worksheet Budget-Friendly Options
  6. Navigating Glencoe Covalent Bonding Chapter 8 Worksheet eBook Formats
    - ePub, PDF, MOBI, and More
    - Glencoe Covalent Bonding Chapter 8 Worksheet Compatibility with Devices
    - Glencoe Covalent Bonding Chapter 8 Worksheet Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Glencoe Covalent Bonding Chapter 8 Worksheet
    - Highlighting and Note-Taking Glencoe Covalent Bonding Chapter 8 Worksheet
    - Interactive Elements Glencoe Covalent Bonding Chapter 8 Worksheet
  8. Staying Engaged with Glencoe Covalent Bonding Chapter 8 Worksheet
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Glencoe Covalent Bonding Chapter 8 Worksheet
  9. Balancing eBooks and Physical Books Glencoe Covalent Bonding Chapter 8 Worksheet
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Glencoe Covalent Bonding Chapter 8 Worksheet
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Glencoe Covalent Bonding Chapter 8 Worksheet
    - Setting Reading Goals Glencoe Covalent Bonding Chapter 8 Worksheet
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Glencoe Covalent Bonding Chapter 8 Worksheet

- Fact-Checking eBook Content of Glencoe Covalent Bonding Chapter 8 Worksheet
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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