

Activity C:
Calculating
specific heat

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Introduction: The specific heat capacity of a substance is the amount of energy needed to change the temperature of that substance by 1 °C. Specific heat capacity can be calculated using the following equation:
$$q = mc\Delta T$$

In the equation q represents the amount of heat energy gained or lost (in **joules**), m is the mass of the substance (in grams), c is the specific heat capacity of the substance (in J/g °C), and ΔT is the temperature change of the substance (in °C).

Goal: Calculate the specific heat capacities of copper, granite, lead, and ice.

- Solve:** When you mix two substances, the heat gained by one substance is equal to the heat lost by the other substance. Suppose you place 125 g of aluminum in a calorimeter with 1,000 g of water. The water changes temperature by 2 °C and the aluminum changes temperature by -74.95 °C.

- Water has a known specific heat capacity of 4.184 J/g °C. Use the specific heat equation to find out how much heat energy the water gained (q).

$$\Delta T Q = 1000 \times 4.18 \times 2Q = 8360 \text{ J} \quad 8360 \text{ J}; 8.4 \text{ kJ}$$

- Assume that the heat energy gained by the water is equal to the heat energy lost by the aluminum. Use the specific heat equation to solve for the specific heat of aluminum. (Hint: Because heat energy is lost, the value of q is negative.)

0.893

Aluminum's accepted specific heat value is 0.900 J/g °C. Use this value to check your work.

- Calculate:** Use the Gizmo to mix 200 g of copper at 100 °C with 1,000 g of water at 20 °C.

- What is the final temperature? 21.42
- Calculate the temperature change of each substance by subtracting the initial temperature from the final temperature.

ΔT_{water} 1.49

ΔT_{copper} 20

- How much heat energy (q) did the water gain? 18
- Now solve for the specific heat (c) of copper; it is 20°C

Gizmo Calorimetry Lab Answers Key

Yicheng Fang

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Gizmo Calorimetry Lab Answers Key

1. Understanding the eBook Gizmo Calorimetry Lab Answers Key
 - The Rise of Digital Reading Gizmo Calorimetry Lab Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Gizmo Calorimetry Lab Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gizmo Calorimetry Lab Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gizmo Calorimetry Lab Answers Key
 - Personalized Recommendations
 - Gizmo Calorimetry Lab Answers Key User Reviews and Ratings
 - Gizmo Calorimetry Lab Answers Key and Bestseller Lists
5. Accessing Gizmo Calorimetry Lab Answers Key Free and Paid eBooks

- Gizmo Calorimetry Lab Answers Key Public Domain eBooks
 - Gizmo Calorimetry Lab Answers Key eBook Subscription Services
 - Gizmo Calorimetry Lab Answers Key Budget-Friendly Options
6. Navigating Gizmo Calorimetry Lab Answers Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Gizmo Calorimetry Lab Answers Key Compatibility with Devices
 - Gizmo Calorimetry Lab Answers Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gizmo Calorimetry Lab Answers Key
 - Highlighting and Note-Taking Gizmo Calorimetry Lab Answers Key
 - Interactive Elements Gizmo Calorimetry Lab Answers Key
 8. Staying Engaged with Gizmo Calorimetry Lab Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gizmo Calorimetry Lab Answers Key
 9. Balancing eBooks and Physical Books Gizmo Calorimetry Lab Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gizmo Calorimetry Lab Answers Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gizmo Calorimetry Lab Answers Key
 - Setting Reading Goals Gizmo Calorimetry Lab Answers Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gizmo Calorimetry Lab Answers Key
 - Fact-Checking eBook Content of Gizmo Calorimetry Lab Answers Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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