



Name:

XXXXXX

Date:

XXXXXX

Student Exploration: Stoichiometry

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL)

bottle of orange juice?

176g

2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150

cookies?

15 sticks

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The **Stoichiometry Gizmo** allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe₂O₃

1

CO

3

Fe

2

CO₂

3

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

Reactants

Iron:

2

Oxygen:

6

Carbon:

3

Products

Iron:

2

Oxygen:

6

Carbon:

3

Based on these values, is the equation balanced?

Yes

New equation



New question

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III) oxide (Fe₂O₃)?

Gizmo Student Exploration Stoichiometry Answer Key

RD Boyd



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