

Name: _____

Date: _____

Student Exploration: Covalent Bonds

Vocabulary: covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. There are eight markers in a full set, but Flora and Frank each only have seven markers. Flora is missing the red marker, and Frank is missing the blue marker.

What can they do so that each has a full set of markers? **They can share the markers.**

2. Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set?

They can give the ones they need to each other.

Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the *Covalent Bonds* Gizmo, atoms form bonds in this way.

To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** (▶) to see the electrons orbiting the nucleus of each atom.



1. The outermost electrons in each atom are called **valence electrons**. How many valence electrons does each fluorine atom have? _____7_____
2. Click **Pause** (⏸). Drag an electron from the left atom to the right atom. Click **Play**.
What happens? **They share an electron.**
3. Click **Pause**, drag an electron from the right atom to the left, and then click **Play**.
What happens now? **They share another one.**

Gizmo Answers Covalent Bonds

Lealyn Burr Clapp

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

Gizmo Answers Covalent Bonds:

Chemical Bonds Phillip Manning, 2009 Modern life is made up of a mind boggling array of materials A simple drinking cup for example might be made of Styrofoam paper or glass depending on the drinkers needs at the moment Home storage cabinets can be made of metal wood or plastic Space shuttles are assembled from silicon steel and hundreds of other materials All of these items owe their properties to the chemical bonds between the atoms that make up the substance Chemical Bonds examines the nature of the chemical bonds answering fundamental questions about how they form how they are broken and how they help define life as we know it

Covalent Bond 35 Success Secrets - 35 Most Asked

Questions on Covalent Bond - What You Need to Know Shirley Soto, 2014-10-14 A Blue Ribbon Covalent bond Guide A covalent bond is a biochemical bond that includes the parting of negatron matches amid particles The steady level of alluring and hideous drives amid particles once they share electrons is recognized like covalent joining For numerous particles the parting of electrons permits every one particle to attain the equal of a complete outside shell comparable to a steady microelectronic arrangement There has never been a Covalent bond Guide like this It contains 35 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Covalent bond A quick look inside of some of the subjects covered Chemical bonding Covalent bond Noncovalent bonding Drug Design Covalent bond History Noncovalent bonding Cation Anion Noncovalent bonding Hydrophobic effect Polar covalent bond Polarity of bonds Noncovalent bonding Boiling Points of Liquids Noncovalent bonding London Dispersion Forces Coordinate Covalent Bond Examples Noncovalent bonding effects Noncovalent bonding H bonding Noncovalent bonding Polar Polar covalent bond Polar molecules Covalent bonds Polarity of covalent bonds Polar covalent bond Polarity of molecules Noncovalent bonding Interaction Noncovalent bonding Van der Waals Forces Covalent bonds Subdivision of covalent bonds Covalent Bond Classification Polar covalent bond Hybrids Noncovalent bonding Electrostatic Interactions Polar covalent bond Nonpolar molecules Covalent Bond Classification Other uses Covalent bonds History Noncovalent bonding Dipole Dipole Noncovalent bonding Protein Folding Structure and much more

The Covalent Bond Henry Sinclair Pickering, 1977

Molecules and the Chemical Bond

Henry A. Bent, 2013-09-30 MOLECULES and the Chemical Bond is about understanding Schrödingers equation for chemical systems In his famous Lectures on Physics Richard Feynman quotes Paul Dirac on what it means to understand an equation I understand what an equation means said Dirac if I have a way of figuring out the characteristics of its solutions without actually solving it That hits the nail on the head Its precisely what Conceptual Valence Bond Theory does for Schrödingers equation A physical understanding of an equation adds Feynman is a completely unmathematical imprecise and inexact thing but absolutely necessary for a physicist It unfolds in MCB in two stages described by Newton as a stage of Analysis a union of

observations and inductions and a stage of Synthesis use of inductions accepted as first principles to explain observations
The books chief vehicle for creating an intuitive understanding of solutions of Schrdingers equation is the worlds largest and
to the authors knowledge virtually only library of line drawings of exclusive orbital models of chemical species electron
density profiles By focussing attention on fundamental physical principles and by avoiding use of atomic orbitals and thereby
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order in leisurely walks so to speak in the big garden that is valence theory picking bouquets to their liking **Chemistry of
the Covalent Bond** Leallyn Burr Clapp,1957 **Covalent Bonding in Crystals, Molecules, and Polymers** James C.
Phillips,1969 Chemical Bonding and the Geometry of Molecules George E. Ryschkewitsch,1965

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