

- 2 Magnesium carbonate reacts with hydrochloric acid.



- (a) What mass of magnesium carbonate would react with excess hydrochloric acid to produce 240 cm³ of carbon dioxide, measured at room temperature and pressure?

Data: 1 mol of any gas occupies 24.0 dm³ at room temperature and pressure

Molar mass of magnesium carbonate = 84.3 g mol⁻¹

- ☒ A 0.843 g
☐ B 8.43 g
☐ C 84.3 g
☐ D 843 g

$$\begin{aligned} \text{Volume of CO}_2 &= 240 \text{ cm}^3 = 0.24 \text{ dm}^3 \\ \text{moles of CO}_2 &= \frac{\text{Volume in dm}^3}{24} = \frac{0.24}{24} = 0.01 \\ \text{MgCO}_3 &\text{ contains same moles as CO}_2 \\ \text{moles} &= \frac{\text{mass}}{\text{Mr}} ; \text{ mass} = \text{moles} \times \text{Mr} \\ &= 0.01 \times 84.3 \\ &= 0.843 \text{ g} \end{aligned}$$

- (b) What is the minimum mass of magnesium carbonate needed to neutralise 50.0 cm³ of 0.250 mol dm⁻³ hydrochloric acid?

- ☐ A 0.423 g
☒ B 0.527 g
☐ C 1.05 g
☐ D 2.11 g

$$\begin{aligned} \text{Concentration of HCl} &= \frac{\text{moles}}{\text{Vol in dm}^3} \\ \text{moles} &= \text{Conc} \times \text{Vol in dm}^3 \\ &= 0.250 \times 0.05 \\ &= 0.0125 \text{ moles of HCl} \end{aligned}$$

$$\begin{aligned} \text{MgCO}_3 : \text{HCl} \\ 1 : 2 \\ \frac{1}{2} : 1 \\ \frac{1}{2} \times 0.0125 : 0.0125 \\ 6.25 \times 10^{-3} : 0.0125 \\ \downarrow \quad \quad \downarrow \\ \text{MgCO}_3 \quad \quad \text{HCl} \\ \text{moles} \quad \quad \text{moles} \end{aligned}$$

- (c) What would be seen at the end of the reaction with excess acid?

- ☒ A A colourless solution
☐ B A coloured solution
☐ C A white precipitate
☐ D A coloured precipitate

all MgCO₃ is utilized
and color of
MgCO₂ colorless

$$\text{mass} = \text{moles} \times \text{Mr}$$

(Total for Question 2 = 3 marks)

$$\begin{aligned} &= 6.25 \times 10^{-3} \times 84.3 \\ &= 0.527 \text{ g} \end{aligned}$$

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Ying-Ying Zheng



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