

## SOLVED QUESTIONS

### Multiple Choice Questions

1.  $K_1$  and  $K_2$  are the velocity constants of forward and backward reactions. The equilibrium constant  $k$  of the reaction is
- (A)  $K_1 \times K_2$  (B)  $K_1 - K_2$   
 (C)  $\frac{K_1}{K_2}$  (D)  $\frac{K_1 + K_2}{K_1 - K_2}$
- Sol.** (C).
2. Which of the following reactions will be favoured at low pressure?
- (A)  $N_2 + O_2 \rightleftharpoons 2NO$  (B)  $H_2 + I_2 \rightleftharpoons 2HI$   
 (C)  $PCl_5 \rightleftharpoons 2PCl_3 + Cl_2$  (D)  $N_2 + 3H_2 \rightleftharpoons 2NH_3$
- Sol.** (C).
3. For the reaction,  $A + B \longrightarrow C + D$ , if active mass of A is kept constant and active mass of B is tripled, the rate of reaction will become
- (A) three times (B) six times  
 (C) eight times (D) nine times
- Sol.** (D).
4. The rate of  $\frac{K_p}{K_c}$  for the reaction:
- $$CO(g) + \frac{1}{2}O_2(g) \rightleftharpoons CO_2(g)$$
- is
- (A) 1 (B)  $RT$   
 (C)  $(RT)^{1/2}$  (D)  $(RT)^{3/2}$
- Sol.** (D).
5. Consider the reaction,
- $$CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$$
- in a closed container at equilibrium. What would be the effect of addition of  $CaCO_3$  on the equilibrium concentration of  $CO_2$ ?
- (A) Increases (B) Unpredictable  
 (C) Decreases (D) Remains unaffected
- Sol.** (D).
6. For the reaction,  $PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$ , the forward reaction at constant temperature is favoured by
- (A) introducing an inert gas at constant volume  
 (B) introducing  $PCl_3(g)$  at constant volume  
 (C) introducing  $PCl_5(g)$  at constant volume  
 (D) introducing  $Cl_2(g)$  at constant volume
- Sol.** (C).
7. A liquid is in equilibrium with its vapour at its boiling point. On an average, the molecules in the two phases have equal
- (A) intermolecular forces  
 (B) potential energy  
 (C) kinetic energy  
 (D) both kinetic energy and potential energy
- Sol.** (C).
8. The unit of equilibrium constant for the reversible reaction

# Multiple Choice Questions In Chemical Equilibrium

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