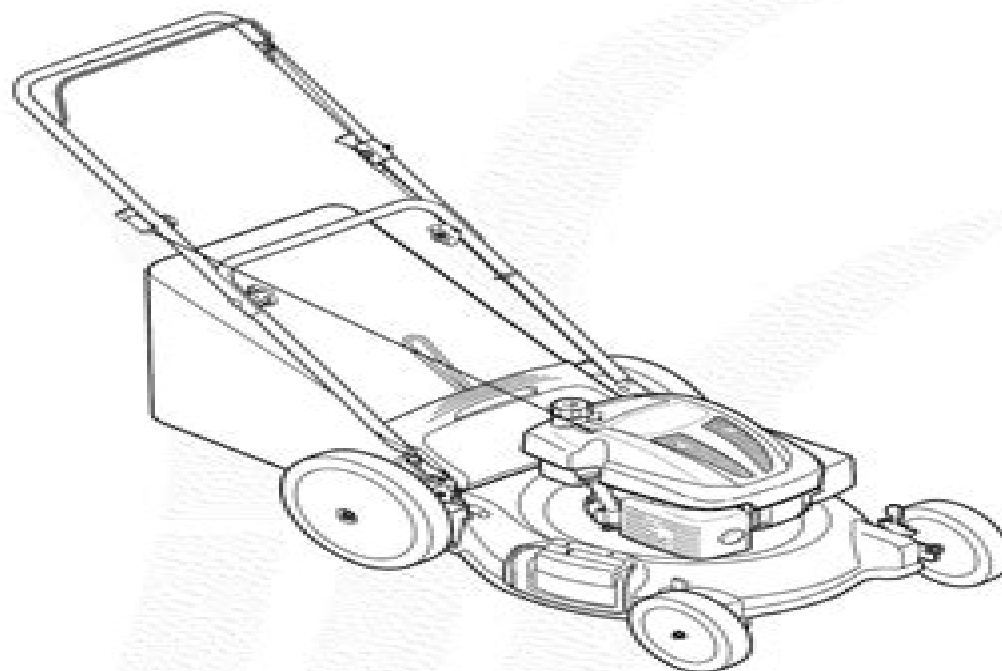


# OPERATOR'S MANUAL



## 21-inch Rotary Lawn Mower — Model 54M

### IMPORTANT

#### READ SAFETY RULES AND INSTRUCTIONS CAREFULLY BEFORE OPERATION

**Warning:** This unit is equipped with an internal combustion engine and should not be used on or near any unimproved forest-covered, brush-covered or grass-covered land unless the engine's exhaust system is equipped with a spark arrester meeting applicable local or state laws (if any). If a spark arrester is used, it should be maintained in effective working order by the operator. In the State of California the above is required by law (Section 4442 of the California Public Resources Code). Other states may have similar laws. Federal laws apply on federal lands. A spark arrester for the muffler is available through your nearest engine authorized service dealer or contact the service department, P.O. Box 361131 Cleveland, Ohio 44136-0019.

# Mtd Lawn Manual

**Rachel Sandford**



## **Mtd Lawn Manual:**

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 identify the oxidation state of each element in the following a na  $2\text{SO}_4$  b  $\text{MnO}_4^-$  c  $\text{KClO}_3$  d  $\text{CH}_3\text{OH}$  2 identify the oxidation state of nitrogen in the following a  $\text{NO}_3^-$  b  $\text{NO}_2$  c  $\text{NO}$

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you  $23 \text{ Fe} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}^{2+}$   
oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$  reduction  $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$   $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3 \text{ Ag}$  oxidation reduction  $25 \text{ Cu} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$  oxidation reduction  $26 \text{ O}_2$

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