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> Teacher's resource answers

Chapter 16

1 Add aqueous silver nitrate solution, [1]

observe a pale cream precipitate [1]

which is partially soluble in ammonia solution, suggests bromide ions. [1]

$\text{Ag}^+(\text{aq}) + \text{Br}^-(\text{aq}) \rightarrow \text{AgBr}(\text{s})$ [1]

2 Number of moles

use $PV = nRT$, rearranged to

give $n = \frac{PV}{RT}$ [1]

$$n = \frac{1.01 \times 10^5 \times 61 \times 10^{-4}}{8.314 \times 373} \quad [1]$$

$$= 1.99 \times 10^{-3} \text{ mol} \quad [1]$$

molar mass

$$= \frac{0.218}{1.99 \times 10^{-3}} = 109.6 \text{ g mol}^{-1} \quad [1]$$

The relative molecular mass of $\text{CH}_3\text{CH}_2\text{Br}$ is 109, which is close to the value above. [1]

Note the need to convert the temperature to an absolute temperature.

3 a $\text{CH}_3\text{CH}_2\text{Br} + \text{KCN} \rightarrow \text{CH}_3\text{CH}_2\text{CN} + \text{KBr}$ [1]

b The substitution of the Br by CN lengthens the carbon chain. [1]

c



curly arrow from lone pair or negative charge on cyanide ion [1]

dipoles on the C-Br bond [1]

curly arrow from the C-Br bond onto the bromine [1]

Mastering Chemistry Answers Chapter 16

Julia Schneider



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