

KINETICS Practice Problems and Solutions

Name: _____
 Period: _____ Date: _____

AP Chemistry
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the $[H_2]$, which will increase the rate, but has no effect on k .
 - Due to Arrhenius equation, changing temperature changes the value of k .
 - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k .

6. Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$, mol/L	$[H_2]$, mol/L	Initial Rate, mol/(L · s)
0.0110	0.0075	3.20×10^{-4}
0.0110	0.0150	6.40×10^{-4}
0.0220	0.0150	6.39×10^{-4}

- What is the order with respect to:
 P_4 0
 H_2 1
 - Write the rate law for this reaction. $rate = k[H_2]$
 - Determine the value and units of the rate constant, k . plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 0.0427 s^{-1}$
7. Consider the reaction: $SO_2 + O_3 \rightarrow SO_3 + O_2$. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$, mol/L	$[O_3]$, mol/L	Initial Rate, mol/(L · s)
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to:
 SO_2 2
 O_3 0
 - Write the rate law for this reaction. $rate = k[SO_2]^2[O_3]^0$
 - Determine the value and units of the rate constant, k . plug and chug using the rate law & data from exp't 1 and solving for k , we get $k = 2.36 mol/L^{-2} s^{-1}$
8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$
- Write the overall balanced chemical equation. $2 A_2 + B_2 \rightarrow 2 C$
 - Identify any intermediates within the mechanism. R
 - What is the order with respect to each reactant? $A_2: 1^{st}$; $B_2: 1^{st}$

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Table of Contents Mastering Chemistry Answer Key Kinetics

1. Understanding the eBook Mastering Chemistry Answer Key Kinetics
 - The Rise of Digital Reading Mastering Chemistry Answer Key Kinetics
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Answer Key Kinetics
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Answer Key Kinetics
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Answer Key Kinetics
 - Personalized Recommendations

- Mastering Chemistry Answer Key Kinetics User Reviews and Ratings
- Mastering Chemistry Answer Key Kinetics and Bestseller Lists
- 5. Accessing Mastering Chemistry Answer Key Kinetics Free and Paid eBooks
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 - Mastering Chemistry Answer Key Kinetics Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Answer Key Kinetics
 - Highlighting and Note-Taking Mastering Chemistry Answer Key Kinetics
 - Interactive Elements Mastering Chemistry Answer Key Kinetics
- 8. Staying Engaged with Mastering Chemistry Answer Key Kinetics
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Answer Key Kinetics
- 9. Balancing eBooks and Physical Books Mastering Chemistry Answer Key Kinetics
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Answer Key Kinetics
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mastering Chemistry Answer Key Kinetics
 - Setting Reading Goals Mastering Chemistry Answer Key Kinetics
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastering Chemistry Answer Key Kinetics
 - Fact-Checking eBook Content of Mastering Chemistry Answer Key Kinetics

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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