

Go to the PHET site by searching for "phet Gases Intro". Run the HTML simulator by pressing the Play button. Answer the following questions. Here is the link: https://phet.colorado.edu/en/simulations/gases-intro?sim_gases-intro_en.html

Directions for INTRO-

1. After opening the website above, click on the **INTRO** box.
2. To add a gas to the container, move the air pump handle up and down, three or four times.
 - a. After each of the following changes, describe what happens to the air pressure reading, that you see in the gauge in the upper right corner of the container.
 - i. Move the toggle switch on the "Heat/Cool" bucket below the container UP, indicating an increase in temperature.
As the temp increases, the molecules should move more quickly, colliding more often and **INCREASING THE PRESSURE**
 - ii. Predict what would happen if you had cooled the temp.
As the temp decreases, the molecules should move more slowly, colliding less often and **DECREASING THE PRESSURE**
 - iii. Now move the handle at the top of the container to the left, briefly, and close again. You have allowed some gas to escape.
As the # of gas particles decreases, the molecules should collide less often and **DECREASE THE PRESSURE**
 - iv. Now move the handle, to the left of the container, towards the right. After you describe the change in the pressure reading, what did you just do to the volume of the container?
As the volume decreases, the molecules should collide more often and **INCREASE THE PRESSURE**
 - v. Predict what would happen to the pressure if you increase the volume of the container. Try it.
As the volume increases, the molecules should collide less often and **DECREASE THE PRESSURE**
3. Now move your cursor to the bottom of the page and click on the **LAWS** box. Once open, move the air pump handle up and down again to add gas to your container.
4. To simulate one of the gas laws, in the box off to the right that says **HOLD CONSTANT**, click on only **temperature**.
 - a. What is the pressure in your container? **[answers may vary]**
 - b. At this setting, with the temp constant, and without changing the number of gas molecules in the container, what two variables could be changed?
PRESSURE & VOLUME
 - c. What gas law focuses on those two variables?
BOYLE'S LAW
 - d. Make a change to one of these variables and identify the resulting impact on the other variable. **[answers may vary]**

Intro To Gases Quest Answer Key

RD Boyd



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