

Intro to Acids and Bases

Name: _____
 Class: _____
 Date: _____

Define each term: Definitions of acids and bases:

Arrhenius Definitions:

- acid substances that release dissolved hydrogen ions (H^+) (acids H^+ gases formed in gases as the hydrogens form H_2 and H_2O (acid H_2O) and conjugated anions)
- base substances which release hydroxide ions (OH^-)

Brønsted-Lowry Definitions:

- acid substances that donate a proton (H^+) (acid donate)
- base substances which accept a proton (H^+) (base receive)

Table 1: Representative reactions for H^+ (acid) react with base representative reactions:

Reaction	Eq.
1. $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$	$H^+(aq) + OH^-(aq)$
2. $HNO_3(aq) + NaOH(aq) \rightarrow NaNO_3(aq) + H_2O(l)$	$H^+(aq) + OH^-(aq)$
3. $H_2SO_4(aq) + NaOH(aq) \rightarrow NaHSO_4(aq) + H_2O(l)$	$H^+(aq) + OH^-(aq)$
4. $H_2SO_4(aq) + NaOH(aq) \rightarrow Na_2SO_4(aq) + 2H_2O(l)$	$2H^+(aq) + 2OH^-(aq)$
5. $HCl(g) + NaOH(s) \rightarrow NaCl(s) + H_2O(l)$	$H^+(g) + OH^-(s)$

Critical Thinking Questions

- Are all of the reactions in Table 1 that are representative acids?
 (a) no all some none
- Are all of the products in Table 1 that are representative bases?

Class work

- Are all of the products in Table 1 that are Brønsted-Lowry acids?

Individual questions for students:

- Are all of the products in Table 1 that are Brønsted-Lowry bases?

Group questions:

- Are products that are not an acid and base? Explain.

Discussion and class work:

- Class members write on their own sheets how they define different conditions? Explain.

Homework

- Write your answers: "Are substances that are changing acids and bases Brønsted-Lowry acids?" Explain.

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William B. Jensen



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