

Intro to Acids and Bases

Name: _____
 Class: _____
 Date: _____

Define each term: Definitions of acids and bases:

Arrhenius-definitions:

- acid substances that release dissolved cations hydrogen (H^+) (actual H^+ seldom formed in water as the hydronium ion, H_3O^+ , and H^+ is used to describe the proton)
- base substances which release hydroxide ions (OH^-)

Brønsted-Lowry definitions:

- acid substances that donate a proton (H^+) (also accepted)
- base substances which accept a proton (H^+) (also donated)

Table 1: Representative reactions for H^+ (also known as the Lewis acid) and bases according to Arrhenius:

Reaction	Eq.
1. $HCl(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + Cl^-(aq)$	$H^+(aq) + H_2O(l)$
2. $HNO_3 + H_2O \rightleftharpoons H_3O^+ + NO_3^-$ (strong acid)	$H^+ + H_2O(l)$
3. $HNO_2(aq) + H_2O \rightleftharpoons H_3O^+ + NO_2^-$ (weak acid)	$H^+(aq) + H_2O$
4. $H_2SO_4 + H_2O \rightleftharpoons H_3O^+ + HSO_4^-$ (strong)	$2 H^+(aq) + H_2O$
5. $H_2SO_4 + H_2O \rightleftharpoons H_3O^+ + HSO_4^-$ (weak)	$H^+ + H_2O$

Critical Thinking Questions

1. List acid characteristics for Table 1 that are Arrhenius acids:

aq, H^+ , H_3O^+ , Cl^- , NO_3^-

2. List acid characteristics for Table 1 that are Arrhenius bases:

OH^-

3. List acid characteristics in Table 1 that are Arrhenius acids and bases:

H_2SO_4 , HNO_3 , HNO_2 , H_2SO_4 , H_2SO_4 , HSO_4^-

4. List acid characteristics in Table 1 that are Arrhenius acids and bases:

H_2SO_4 , HNO_3 , HNO_2

5. Is it possible for an acid to be an Arrhenius base? Explain:

Yes, because H_2SO_4 , HNO_3 , HNO_2 can be bases.

6. List acid characteristics in Table 1 that are Arrhenius acids and bases:

aq, H^+ , H_3O^+ , Cl^- , NO_3^- , HNO_2 , H_2SO_4 , HSO_4^- , H_2SO_4 , H_2SO_4 , HSO_4^-

7. Is it possible for an acid to be an Arrhenius base? Explain:

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Chemquest 48

Raffaela Di Napoli



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