

Virtual Lab: Empirical Formula of a Hydrate

Background:

Hydrates are ionic compounds (salts) that have a definite amount of water (water of hydration) as part of their structure. The water is chemically combined with the salt in a **definite ratio**. Ratios vary in different hydrates but are specific for any given hydrate.

The formula of a hydrate is represented in a special manner. The hydrate of copper sulfate in this experiment is listed below:



In the formula, the unit formula for the salt appears first, and the water formula is last. *The raised dot means that the water is loosely bonded to the salt.* The coefficient x stands for the number of molecules of water bonded to one unit of salt. This special formula, like all other formulas, illustrates the law of definite composition.

When hydrates are heated, the "water of hydration" is released as vapor. The remaining solid is known as the **anhydrous salt**. The general reaction for heating a hydrate is:



Some anhydrous salts can absorb moisture from the air to become hydrated. A common example is silica gel, made from sodium silicate which is usually packaged with optical or electronic devices shipped by boat. The silica gel protects the devices from high humidity.



We can find the percent of water in a hydrate experimentally by accurately determining the mass of the hydrate and the mass of the anhydrous salt. The difference in mass is due to the water lost by the hydrate. The percentage of water in the original hydrate can easily be calculated:

$$\text{Percent H}_2\text{O} = \frac{\text{mass of H}_2\text{O}}{\text{mass of hydrate}} \times 100$$

Hydrates Lab Answers

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A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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