

24. Give an example to show oxidation and reduction go simultaneously. Balance the following equation by oxidation number or ion electron method
 $\text{HNO}_3 + \text{H}_2\text{S} \rightarrow \text{SO}_2 + \text{NO}_2 + \text{H}_2\text{O}$ 2+3
25. How would you detect the presence of nitrogen in organic compound by Lassaigne test method. 5
26. Deduce the relationship between molecular mass and vapour density of a gas using Avogadro's hypothesis. 3+2
 An Oxide of nitrogen contains half of its own volume of nitrogen and has vapour density equal to 15. Determine molecular formula of the nitrogen oxide.
27. Why can't HBr prepared by using NaBr and Conc. H_2SO_4 ? Give a balanced chemical reaction for the preparation of HCl. How would you test presence of the Cl^- or Br^- ions in its aqueous solution? 2+1+2
28. Describe the Kipp's apparatus method for the manufacture of H_2S . Give a reaction to show H_2S is a reducing agent. 4+1
29. Write short notes on: 2.5+2.5
 i) Hydrometallurgy
 ii) Zone of refining.

Group 'C'

Attempt any two questions.

2x10=20

30. How is ammonia manufactured by Haber's process? What happens when ammonia is passed through
 i) Copper sulphate solution
 ii) heated copper oxide.
 Give reactions for the conversion of ammonia to nitric acid. 6+2+2
31. Write the principle and process along with self-explanatory diagram for the manufacture of sodium by Down's method. Why is sodium fire at laboratory not extinguished by adding water? Mention any two advantages of Down's process. How is sodium converted into washing soda. 6+1+1+2

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J Ma



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