

Chemistry

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(Chapter 6) (General Principle and Processes of Isolation of Elements) (Exercises)

(Class 12)

is filtered, dried, and heated to get back pure alumina.

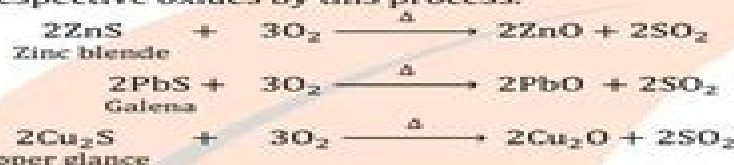


Question 6.14:

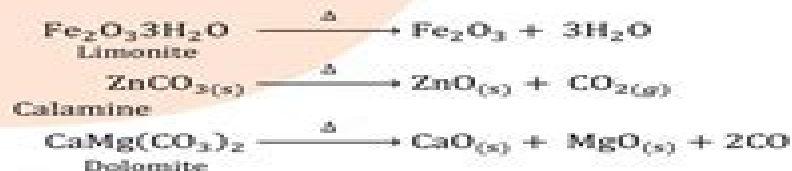
Giving examples, differentiate between 'roasting' and 'calcination'.

Answer 6.14:

Roasting is the process of converting sulphide ores to oxides by heating the ores in a regular supply of air at a temperature below the melting point of the metal. For example, sulphide ores of Zn, Pb, and Cu are converted to their respective oxides by this process.



On the other hand, calcination is the process of converting hydroxide and carbonate ores to oxides by heating the ores either in the absence or in a limited supply of air at a temperature below the melting point of the metal. This process causes the escaping of volatile matter leaving behind the metal oxide. For example, hydroxide of Fe, carbonates of Zn, Ca, Mg are converted to their respective oxides by this process.



Question 6.15:

How is 'cast iron' different from 'pig iron'?

Answer 6.15:

The iron obtained from blast furnaces is known as pig iron. It contains around 4% carbon and many impurities such as S, P, Si, Mn in smaller amounts.

Cast iron is obtained by melting pig iron and coke using a hot air blast. It contains a lower amount of carbon (3%) than pig iron. Unlike pig iron, cast iron is extremely hard and brittle.

Question 6.16:

Differentiate between "minerals" and "ores".

Answer 6.16:

> **Minerals:** Minerals are naturally occurring chemical substances containing metals. They are found in the Earth's crust and are obtained by mining.

> **Ores:** Ores are rocks and minerals viable to be used as a source of metal.

For example, there are many minerals containing zinc, but zinc cannot be extracted profitably (conveniently and economically) from all these minerals.

Zinc can be obtained from zinc blende (ZnS), calamine (ZnCO_3), Zincite (ZnO) etc.

Note: All ores are minerals but all minerals are not ores.

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