

Physics

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(Chapter 13) (Kinetic Theory) (Additional Exercises)
(Class 11)

Question 13.14:

Given below are densities of some solids and liquids. Give rough estimates of the size of their atoms:

Substance	Atomic Mass (u)	Density (10^3 Kg m^{-3})
Carbon (diamond)	12.01	2.22
Gold	197.00	19.32
Nitrogen (liquid)	14.01	1.00
Lithium	6.94	0.53
Fluorine (liquid)	19.00	1.14

[Hint: Assume the atoms to be 'tightly packed' in a solid or liquid phase, and use the known value of Avogadro's number. You should, however, not take the actual numbers you obtain for various atomic sizes too literally. Because of the crudeness of the tight packing approximation, the results only indicate that atomic sizes are in the range of a few Å].

Answer 13.14:

Substance	Radius (Å)
Carbon (diamond)	1.29
Gold	1.59
Nitrogen (liquid)	1.77
Lithium	1.73
Fluorine (liquid)	1.88

Atomic mass of a substance = M and Density of the substance = ρ

Avogadro's number = $N = 6.023 \times 10^{23}$

Volume of N number of molecules = $\frac{4}{3} \pi r^3 N$... (i)

Volume of one mole of a substance = $\frac{M}{\rho}$... (ii)

$$\frac{4}{3} \pi r^3 N = \frac{M}{\rho} \Rightarrow r = \sqrt[3]{\frac{3M}{4\pi\rho N}}$$

For gold: $M = 197.00 \times 10^{-3} \text{ kg}$ $\rho = 19.32 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 197 \times 10^{-3}}{4\pi \times 19.32 \times 10^3 \times 6.23 \times 10^{23}} \right)^{\frac{1}{3}} = 1.77 \text{ Å}$$

Hence, the radius of a liquid nitrogen atom is 1.77 Å.

For lithium: $M = 6.94 \times 10^{-3} \text{ kg}$ $\rho = 0.53 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 6.94 \times 10^{-3}}{4\pi \times 0.53 \times 10^3 \times 6.23 \times 10^{23}} \right)^{\frac{1}{3}} = 1.73 \text{ Å}$$

Hence, the radius of a lithium atom is 1.73 Å.

For liquid fluorine: $M = 19.00 \times 10^{-3} \text{ kg}$ and $\rho = 1.14 \times 10^3 \text{ kg m}^{-3}$

$$\therefore r = \left(\frac{3 \times 19 \times 10^{-3}}{4\pi \times 1.14 \times 10^3 \times 6.023 \times 10^{23}} \right)^{\frac{1}{3}} = 1.88 \text{ Å}$$

Hence, the radius of a liquid fluorine atom is 1.88 Å.

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Table of Contents Mastering Physics Chapter 13 Solutions

1. Understanding the eBook Mastering Physics Chapter 13 Solutions
 - The Rise of Digital Reading Mastering Physics Chapter 13 Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Physics Chapter 13 Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Physics Chapter 13 Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Physics Chapter 13 Solutions
 - Personalized Recommendations
 - Mastering Physics Chapter 13 Solutions User Reviews and Ratings
 - Mastering Physics Chapter 13 Solutions and Bestseller Lists

5. Accessing Mastering Physics Chapter 13 Solutions Free and Paid eBooks
 - Mastering Physics Chapter 13 Solutions Public Domain eBooks
 - Mastering Physics Chapter 13 Solutions eBook Subscription Services
 - Mastering Physics Chapter 13 Solutions Budget-Friendly Options
6. Navigating Mastering Physics Chapter 13 Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Mastering Physics Chapter 13 Solutions Compatibility with Devices
 - Mastering Physics Chapter 13 Solutions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Physics Chapter 13 Solutions
 - Highlighting and Note-Taking Mastering Physics Chapter 13 Solutions
 - Interactive Elements Mastering Physics Chapter 13 Solutions
8. Staying Engaged with Mastering Physics Chapter 13 Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Physics Chapter 13 Solutions
9. Balancing eBooks and Physical Books Mastering Physics Chapter 13 Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Physics Chapter 13 Solutions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mastering Physics Chapter 13 Solutions
 - Setting Reading Goals Mastering Physics Chapter 13 Solutions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mastering Physics Chapter 13 Solutions
 - Fact-Checking eBook Content of Mastering Physics Chapter 13 Solutions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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