

Chapter 4: Supramolecular Problems - Answers

1. Assign the oxidation number to each atom in the following elements, molecules and ionic substances:

- a) CO_2 b) CH_4 c) HNO_3 d) Ca(OH)_2 e) H_2SO_4 f) Ag
 $\text{C} = +4$ $\text{C} = +4$ $\text{N} = +5$ $\text{Ca} = +2$ $\text{S} = +6$ $\text{Ag} = 0$
 $\text{O} = -2$ $\text{H} = +1$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{Ag} = 0$
- g) Fe h) BaCl_2 i) CO_2 j) H_2O k) HNO_3 l) Fe
 $\text{Fe} = 0$ $\text{Ba} = +2$ $\text{C} = +4$ $\text{H} = +1$ $\text{N} = +5$ $\text{Fe} = 0$
 $\text{Cl} = -1$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{Fe} = 0$

2. Assign the oxidation number to each atom in the following ions:

- a) NO^+ b) CO_3^{2-} c) HNO_2 d) SO_4^{2-} e) PO_4^{3-} f) HCO_3^-
 $\text{N} = +2$ $\text{C} = +3$ $\text{N} = +3$ $\text{S} = +6$ $\text{P} = +5$ $\text{C} = +3$
 $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$

3. For the following reactions identify if it is an oxidation-reduction reaction. If it is a redox reaction identify the oxidizing agent, the reducing agent, the substance being oxidized and the substance being reduced. Identify the oxidation number change to help to support these observations:



Redox rxn. Ca oxidized (oxidizing agent); HCl is $\text{H}^+ \rightarrow \text{H}_2$
 HCl reduced (reducing agent); Ca is $0 \rightarrow +2$



Not a redox rxn. Ca is $0 \rightarrow 0$ (not oxidized/reduced)



Redox rxn. H_2 oxidized (oxidizing agent); O is $0 \rightarrow -1$
 H_2 reduced (reducing agent); O is $0 \rightarrow -1$



Not a redox reaction. No change in ox. #



Mastering Chemistry Answer Key Ch 14

Morris Hein, Susan Arena, Cary Willard



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Table of Contents **Mastering Chemistry Answer Key Ch 14**

1. Understanding the eBook **Mastering Chemistry Answer Key Ch 14**
 - The Rise of Digital Reading **Mastering Chemistry Answer Key Ch 14**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Mastering Chemistry Answer Key Ch 14**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Mastering Chemistry Answer Key Ch 14**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **Mastering Chemistry Answer Key Ch 14**
 - Personalized Recommendations
 - **Mastering Chemistry Answer Key Ch 14** User Reviews and Ratings
 - **Mastering Chemistry Answer Key Ch 14** and Bestseller Lists

5. Accessing Mastering Chemistry Answer Key Ch 14 Free and Paid eBooks
 - Mastering Chemistry Answer Key Ch 14 Public Domain eBooks
 - Mastering Chemistry Answer Key Ch 14 eBook Subscription Services
 - Mastering Chemistry Answer Key Ch 14 Budget-Friendly Options
6. Navigating Mastering Chemistry Answer Key Ch 14 eBook Formats
 - ePub, PDF, MOBI, and More
 - Mastering Chemistry Answer Key Ch 14 Compatibility with Devices
 - Mastering Chemistry Answer Key Ch 14 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Chemistry Answer Key Ch 14
 - Highlighting and Note-Taking Mastering Chemistry Answer Key Ch 14
 - Interactive Elements Mastering Chemistry Answer Key Ch 14
8. Staying Engaged with Mastering Chemistry Answer Key Ch 14
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Answer Key Ch 14
9. Balancing eBooks and Physical Books Mastering Chemistry Answer Key Ch 14
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Answer Key Ch 14
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mastering Chemistry Answer Key Ch 14
 - Setting Reading Goals Mastering Chemistry Answer Key Ch 14
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mastering Chemistry Answer Key Ch 14
 - Fact-Checking eBook Content of Mastering Chemistry Answer Key Ch 14
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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