

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                        |                             |                   |
|------------------------|-----------------------------|-------------------|
| 1. autoionization      | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization          | 5. $K_a$                    | 6. $K_b$          |
| 7. ionization constant | 8. $pH$                     |                   |

## Key Equation

- |                       |                       |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the  $K_a$  or  $K_b$  for each acid or base. (For weak acids and bases, write the chemical equation.)

1.  $HNO_3$  is a strong acid.  $K_a =$  \_\_\_\_\_
2.  $HCl$  is a strong acid.  $K_a =$  \_\_\_\_\_
3.  $H_2SO_4$  is a strong acid.  $K_a =$  \_\_\_\_\_
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