

CHAPTER 11 REVIEW

Ions in Aqueous Solutions and Colligative Properties

SECTION REVIEW

IMPORT SKILLS Answer the following questions in the space provided.

1. Identify the ions represented by the signs in these descriptions on the left.

- | | |
|---|-------------------------------------|
| <u> </u> an ionic compound that is quite soluble in water | (a) NaCl |
| <u> </u> an ionic compound that is not very soluble in water | (b) NaNO_3 |
| <u> </u> a molecular compound that contains no oxygen | (c) AgCl |
| <u> </u> a molecular compound that does not contain hydrogen | (d) $\text{C}_2\text{H}_5\text{OH}$ |

2. Consider separately separately-saturated solutions of various liquid solutes in water. In each case, the following statements:

- | | |
|---|---|
| <u> </u> | a. The change in the boiling point decreases with the density of the solid solute, with the increasing solute density causing greater increases. |
| <u> </u> | b. The change in the boiling point decreases with the density of the solid solute, with the increasing solute density causing greater increases. |
| <u> </u> | c. The change in the boiling point becomes greater as the concentration of the solute in solution increases decreases. |

3. a. Explain why compound A is more soluble in water than compound B, which has the formula of a typical molecular substance.

b. Identify any species ions in the system from dissociation process.

c. Write the net ionic equation for the formation of water vapor.

4. Explain why applying road salt helps melt ice on city streets during winter.

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