- 1.16. 1 lb mol = 453.59 g mol (App. 1). At 1 atm, 20° C, Number of molecules of oxygen, from Avogadro's number: $0.209 \times 6.022 \times 10^{23} \times 453.59 = 5.71 \times 10^{23}$ Number is not affected by temperature or pressure. At 1 atm and 30° C, in 1 kg mol the number is: $0.209 \times 6.022 \times 10^{23} \times 1000 = 1.259 \times 10^{24}$
- 1.17. The internal energy E of an object is the product of its mass and the sum of its kinetic energy, potential energy, and enthalpy. The change in internal energy, ΔE, therefore equals the left-hand side of Eq. (1.59), and we can say,

 $\Delta E = Q - W$

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Thomas Griffiths

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