

Name: _____ Period: _____

Temperature Conversions and Heat Practice

Please show all work to receive credit

$$K = ^\circ C + 273$$

$$^\circ F = (9/5 \times ^\circ C) + 32$$

$$^\circ C = 5/9(^{\circ}F - 32)$$

Convert the following to Fahrenheit 1) $10^{\circ}C$ _____ 2) $30^{\circ}C$ _____ 3) $0^{\circ}C$ _____	Convert the following to Celsius 4) $32^{\circ}F$ _____ 5) $90^{\circ}F$ _____ 6) $212^{\circ}F$ _____
Convert the following to Kelvin 7) $0^{\circ}C$ _____ 8) $-50^{\circ}C$ _____ 9) $90^{\circ}C$ _____ 10) $-20^{\circ}C$ _____	Convert the following to Celsius 11) $112^{\circ}K$ _____ 12) $200^{\circ}K$ _____ 13) $273^{\circ}K$ _____ 14) $350^{\circ}K$ _____

Use the information below to complete the following problems

$$\text{Heat} = \text{mass} \times \text{specific heat} \times \text{change in temperature}$$

$$\Delta t = \text{final temperature} - \text{initial temperature}$$

$$Q = m c \Delta t$$

$$\Delta t = t_f - t_i$$

Substances	(j/kg•K)	Substance	(j/kg•K)	Substance	(j/kg•K)
Water (liquid)	4186	Copper	385	Iron	449
Steam	1870	Gold	129	Aluminum	897

- 15) How much energy is needed to heat 55 Kg of water from 285 K to 360 K?

- 16) How much energy is released when 55 Kg of steam cools from 558 K to 390 K?

- 17) If 5785 J of energy is used to heat 48 kg of Iron what is it's change in temperature?

- 18) An 75 kg block of metal absorbs 721875 J of heat as the temperature increases from 375 K to 400 K.
What is the specific heat of the metal?

 a. Which metal is it likely to be?
- 19) How much energy is released if 278 kg of Aluminum cools from 698 K to 298 K?

- 20) What is the specific heat of a 3.78 kg object that absorbs 678 J as the temperature increases by 4.25 K?

Het And Its Measurement Answer Key

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nov 6 2021 rhoden who was crowned mr olympia in 2018 at the age of 43 making him the oldest bodybuilder to ever win the title remained committed to bodybuilding until his death while an official

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