



GRADE 11 LIFE SCIENCES MARCH TEST

SECTION A QUESTION 1

1.1 Various options are provided as possible answers to the following questions. Choose the correct answer and write **ONLY** the LETTER (A-D) next to the QUESTION NUMBER (1.1.1-1.1.5) on your answer sheet, for example 1.1.B. D.

1.1.1 Which of the following micro-organisms act as decomposers in ecosystems?

- (1) Fungi
 - (2) Bacteria
 - (3) Plants
 - (4) Fungi
- A 1 and 2
B 1 and 3
C 2 and 4
D 1 and 4

1.1.2 Malaria is spread by:

- A Female *Anopheles* mosquito
- B Tick transmission
- C Direct contact
- D Bite of the tsetse

1.1.3 Protists are unlike bacteria in that they...

- A Have prokaryotic cells
- B Are unicellular
- C Are produced by binary fission
- D Are autotrophic

1.1.4 Which statement best describes the relationship between a parasite and its host?

- A The parasite always harms the host benefit.
- B The host is harmed with the parasite increasing.
- C Both host and parasite benefit.
- D The parasite benefits and the host is unaffected.

1.1.5 The following statements about HIV are NOT true:

- (1) It can be passed to a child from its mother during pregnancy.
 - (2) It can be passed to a child from its mother during breastfeeding.
 - (3) It can be passed from a mother to her child while in the womb.
 - (4) The only transmission of HIV infection is the same as that of other diseases.
- A 1 and 2
B 1 and 3
C 1 only
D 2 and 4

[5x2=10]

1.2 Give the correct **BIOLOGICAL TERM** for each of the following descriptions. Write only the correct **TERM** next to the QUESTION NUMBER (1.2.1-1.2.4) on your answer sheet.

1.2.1 A bacterium which can decompose dead organic matter.

1.2.2 A red blood cell in mammals.

1.2.3 The process of producing a cell which is genetically identical to the parent cell.

1.2.4 Using living organisms to control harmful populations.

[4]

1.3 Indicate whether each of the statements in COLUMN I applies to **A ONLY**, **B ONLY**, **BOTH A and B** or **NONE** of the items in COLUMN II. Write **A ONLY**, **B ONLY**, **BOTH A and B** or **NONE** next to the QUESTION NUMBER (1.3.1-1.3.3).

Column I	Column II
1.3.1 Presence of a slime capsule and well-structured nucleus	<u>Amoeba</u> <u>Paramecium</u>
1.3.2 A disease caused by fungi	<u>Ringworm</u> <u>Aspergillus</u> <u>Black bread mould</u>
1.3.3 Binary fission is the asexual production of...	<u>Amoeba</u> <u>Paramecium</u>

[x2=6]

SECTION A: [16]

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ordering information These 400 entries were reviewed by both educators and scientists to ensure that they are accurate and current and offer students the opportunity to Ask questions and find their own answers Experiment productively Develop patience persistence and confidence in their own ability to solve real problems The entries in the curriculum section are grouped by scientific area Life Science Earth Science Physical Science and Multidisciplinary and Applied Science and by type core materials supplementary materials and science activity books Additionally a section of references for teachers provides annotated listings of books about science and teaching directories and guides to science trade books and magazines that will help teachers enhance their students science education Resources for Teaching Elementary School Science also lists by region and state about 600 science centers museums and zoos where teachers can take students for interactive science experiences Annotations highlight almost 300 facilities that make significant efforts to help teachers Another section describes more than 100 organizations from which teachers can obtain more resources And a section on publishers and suppliers give names and addresses of sources for materials The guide will be invaluable to teachers principals administrators teacher trainers science curriculum specialists and advocates of hands on science teaching and it will be of interest to parent teacher organizations and parents

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web the answer is rather simple if you understand electron configurations the shape of the periodic table mimics the filling of the subshells with electrons let us start with h and he their electron configurations are $1s^1$ and $1s^2$ respectively with he the n = 1 shell is filled

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web photoelectron spectroscopy electron configurations google classroom you might need periodic table using s p d f notation what is the electron configuration for a neutral atom of beryllium choose 1 answer $1s^2 2s^2$ a $1s^2 2s^2 1s^2 2s^1$ b $1s^2 2s^1 2p^1$ c $1s^2 2p^2$ d $1s^2 1p^2$

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web jun 14 2015 this list of electron configurations of elements contains all the elements in increasing order of atomic number to save room the configurations are in noble gas shorthand this means part of the electron configuration has been replaced with the element symbol of the noble gas symbol

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web elements in the second row of the periodic table place their electrons in the 2n shell as well as the 1n shell for instance lithium li has three electrons two fill the 1s orbital and the third is placed in the 2s orbital giving an electron configuration of $1s^2 2s^1$

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web electron configurations and trends answer these discussion questions given the following part of an electron configuration $3s^2$ explain which part refers to the number of electrons the energy level and the sublevel what is the maximum number of electrons that can occupy a single orbital state and explain hunds rule

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group iia

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web the electron removed during the ionization of beryllium $1s^2 2s^2$ is an s electron whereas the electron removed during the ionization of boron $1s^2 2s^2 2p^1$ is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium $1s^2$ ne 10 neon $1s^2 2s^2 2p^6$

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ or simply put $4s^1$ q2 what is the total number of electrons that can be filled in s p and d subshells

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web by building up from hydrogen this table can be used to determine the electron configuration for any atom on the periodic table we will now construct the ground state electron configuration and orbital diagram for a selection of atoms in the first and second periods of the periodic table

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