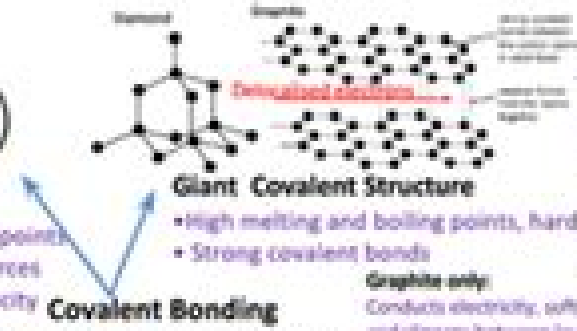
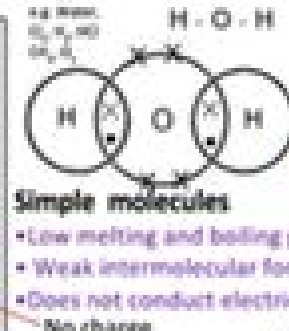




Periodic Table of the Elements

Legend:

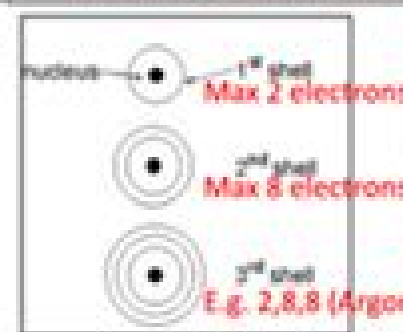
- metals
- alkali metals
- alkali earth metals
- transition metals
- post metals
- metalloids
- non metals
- halogens
- noble gases

An atom:
Number of Protons = Number of Electrons
Mass number = Protons + Neutrons

C2 Chemistry

Periodic table arranges elements by
ATOMIC NUMBER (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
E.g. Lithium 2,1 is a group 1 element.



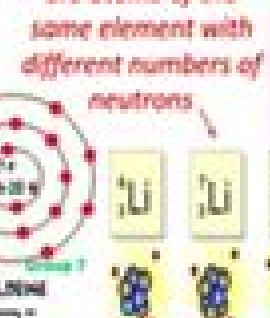
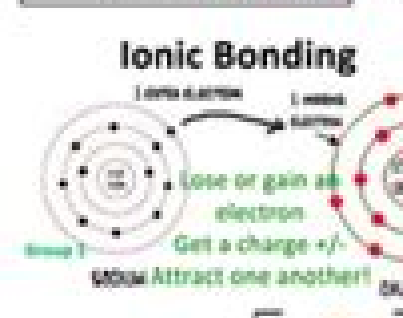
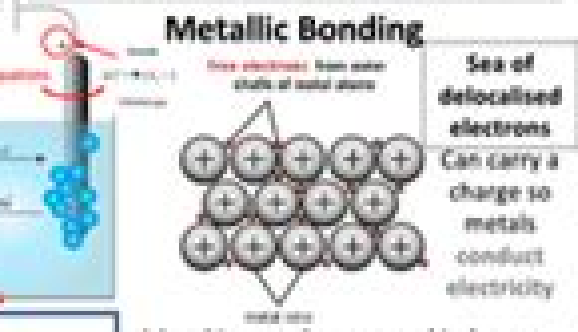
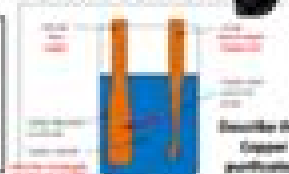
Protons and neutrons both have a mass of 1

ISOTOPES
are atoms of the same element with different numbers of neutrons

E.g. 2,8,8 (Argon)

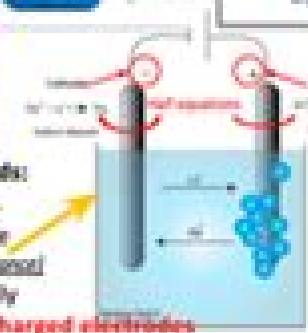
OIL RIG

Oxidation is loss (of electrons)
Reduction is gain (of electrons)



Electrolysis:
Splitting up a substance using electricity

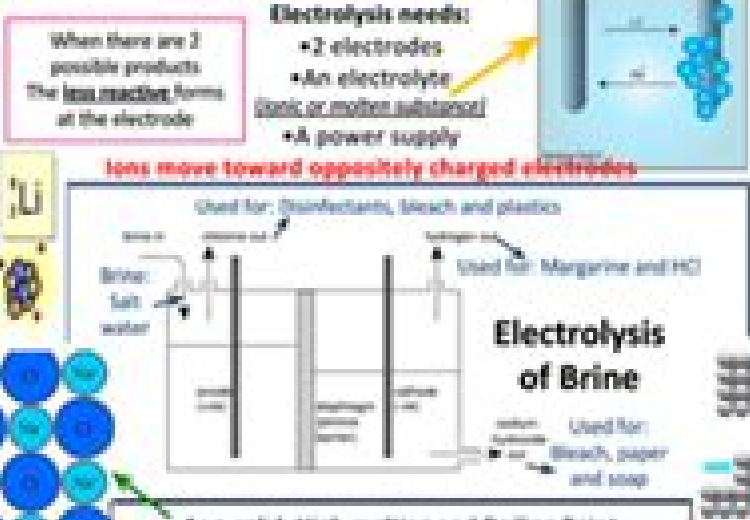
When there are 2 possible products
The less reactive forms at the electrode



Metal ions only arranged in layers
ELECTROSTATIC ATTRACTION
between
+IONS and ELECTRONS

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well



NANO-TECHNOLOGY
Means **REALLY REALLY REALLY** small!!!!

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 - The Rise of Digital Reading Gcse Additional Chemistry C2 Revision Booklet
 - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Additional Chemistry C2 Revision Booklet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gcse Additional Chemistry C2 Revision Booklet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Additional Chemistry C2 Revision Booklet
 - Personalized Recommendations
 - Gcse Additional Chemistry C2 Revision Booklet User Reviews and Ratings
 - Gcse Additional Chemistry C2 Revision Booklet and Bestseller Lists
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 - Gcse Additional Chemistry C2 Revision Booklet Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Gcse Additional Chemistry C2 Revision Booklet
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 - Interactive Elements Gcse Additional Chemistry C2 Revision Booklet
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gcse Additional Chemistry C2 Revision Booklet
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 - Benefits of a Digital Library
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Carving Out Dedicated Reading Time
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 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

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- Interactive and Gamified eBooks

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