

1. How many grams of NaCl are contained in 700. mL of a 0.250 M solution of sodium chloride?

$$0.250 \frac{\text{mol NaCl}}{\text{L}} \times 0.700 \text{ L} \times \frac{58.5 \text{ g NaCl}}{\text{mol}} = 10.2 \text{ g NaCl}$$

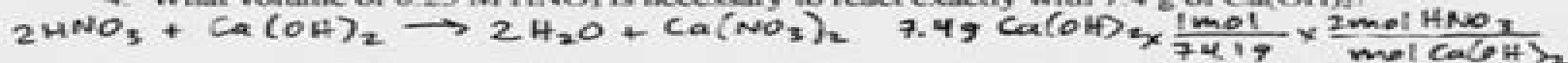
2. The oxidation state of chromium in  $\text{K}_2\text{CrO}_4$  is:

+6

3. The oxidation state of sulfur in the sulfate ion is

+6

4. What volume of 0.25 M  $\text{HNO}_3$  is necessary to react exactly with 7.4 g of  $\text{Ca}(\text{OH})_2$ ?



5. Describe how you would prepare 250. mL of a 0.354 M sodium sulfate solution using solid sodium sulfate crystals.

$$0.354 \frac{\text{mol Na}_2\text{SO}_4}{\text{L}} \times 0.250 \text{ L} \times \frac{142 \text{ g Na}_2\text{SO}_4}{\text{mol}} = 12.6 \text{ g}$$

Transfer 12.6 g  $\text{Na}_2\text{SO}_4$  to a 250 mL vol. flask.

Add distilled water, about 200 mL. Swirl until all solute has dissolved, make up to the mark with distilled water.

$$\times \frac{1 \text{ Liter}}{0.25 \text{ mol HNO}_3} = \boxed{0.80 \text{ L}}$$

- a. What volume, in mL, of 12.1 M hydrochloric acid would be required to make 5.00 L of 2.00 M hydrochloric acid solution?

$$\text{Use } M_1V_1 = M_2V_2$$

$$V_1 = \frac{M_2V_2}{M_1} = \frac{(2.00 \text{ M})(5.00 \text{ L})}{12.1 \text{ M}} = 0.826 \text{ L or } \boxed{826 \text{ mL}}$$

6. Predict the products of the following reactions. Where a reaction occurs, write the net ionic equation. If no reaction occurs, write "N.R."

- a. iron (III) sulfate solution and strontium chloride solution



- b. lead (II) acetate solution and lithium iodide solution



- c. aluminum nitrate solution and calcium bromide solution

N.R.

- d. ammonium hydroxide solution and cobalt (II) nitrate powder



# Gas Variables Worksheet Answers

**Victor M. Corman**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents Gas Variables Worksheet Answers**

1. Understanding the eBook Gas Variables Worksheet Answers
  - The Rise of Digital Reading Gas Variables Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Variables Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Variables Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Variables Worksheet Answers
  - Personalized Recommendations
  - Gas Variables Worksheet Answers User Reviews and Ratings
  - Gas Variables Worksheet Answers and Bestseller Lists
5. Accessing Gas Variables Worksheet Answers Free and Paid eBooks
  - Gas Variables Worksheet Answers Public Domain eBooks
  - Gas Variables Worksheet Answers eBook Subscription Services
  - Gas Variables Worksheet Answers Budget-Friendly Options

6. Navigating Gas Variables Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Variables Worksheet Answers Compatibility with Devices
  - Gas Variables Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Variables Worksheet Answers
  - Highlighting and Note-Taking Gas Variables Worksheet Answers
  - Interactive Elements Gas Variables Worksheet Answers
8. Staying Engaged with Gas Variables Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Variables Worksheet Answers
9. Balancing eBooks and Physical Books Gas Variables Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Variables Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gas Variables Worksheet Answers
  - Setting Reading Goals Gas Variables Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Variables Worksheet Answers
  - Fact-Checking eBook Content of Gas Variables Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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