

13. Complete the following table for the three experiments in Model 2.

	Experiment C	Experiment D	Experiment E
Independent Variable	# of particles	Temp.	Internal External Pressure
Dependent Variable	Volume	Volume	Volume
Controlled Variable(s)	Temp / I+E pressure	External + Internal Pressure / # of particles	Temp / # of particles

14. Provide a molecular level explanation for the increase in volume among the balloons in Experiment C. (How often and/or how hard are the molecules hitting the sides of the container?)

of molecules \uparrow , so number of collisions \uparrow + Volume along with it

collisions \uparrow energy stays the same

15. Provide a molecular level explanation for the increase in volume among the balloons in Experiment D.

Temperature \uparrow increasing the kinetic energy of the particles as well
which \uparrow collisions thus \uparrow pressure

16. Provide a molecular level explanation for the increase in volume among the balloons in Experiment E.

if external forces decrease then it allows internal forces \uparrow
thus \uparrow volume

Gas Variables Pogil Answer

C. E. R. Bruce



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Gas Movements in the Long-period Variables C. E. R. Bruce, 1957
PLC, 1983-01-01

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