

Gas laws practice test

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- ____ 1. Pressure is the force per unit
a. volume. c. length.
b. surface area. d. depth.
- ____ 2. Why does a can collapse when a vacuum pump removes air from the can?
a. The inside and outside forces balance out and crush the can.
b. The unbalanced outside force from atmospheric pressure crushes the can.
c. The atmosphere exerts pressure on the inside of the can and crushes it.
d. The vacuum pump creates a force that crushes the can.
- ____ 3. If the height of mercury in a barometer at 0°C is less than 760 mm Hg, then
a. the atmospheric pressure is less than standard atmospheric pressure.
b. the atmospheric pressure is greater than standard atmospheric pressure.
c. the atmospheric pressure is equal to standard atmospheric pressure.
d. the atmospheric pressure cannot be determined.
- ____ 4. Convert the pressure 0.840 atm to mm Hg.
a. 365 mm Hg c. 638 mm Hg
b. 437 mm Hg d. 780 mm Hg
- ____ 5. Convert the pressure 1.30 atm to kPa.
a. 2 kPa c. 132 kPa
b. 115 kPa d. 245 kPa
- ____ 6. Standard temperature is exactly
a. 100°C. c. 0°C.
b. 273°C. d. 0 K.
- ____ 7. Three samples of gas each exert 740. mm Hg in separate 2 L containers. What pressure do they exert if they are all placed in a single 2 L container?
a. 247 mm Hg c. 1.48×10^3 mm Hg
b. 740 mm Hg d. 2.22×10^3 mm Hg
- ____ 8. If the temperature remains constant, V and P represent the original volume and pressure, and V' and P' represent the new volume and pressure, what is the mathematical expression for Boyle's law?
a. $P'V = V'P$ c. $VP = V'P'$
b. $VV' = PP'$ d. $V' = \frac{VP}{P'}$
- ____ 9. A sample of oxygen occupies 560. mL when the pressure is 800.00 mm Hg. At constant temperature, what volume does the gas occupy when the pressure decreases to 700.0 mm Hg?
a. 80.0 mL c. 600. mL
b. 490. mL d. 640. mL
- ____ 10. If the temperature of a fixed quantity of gas decreases and the pressure remains unchanged,
a. its volume increases. c. its volume decreases.
b. its volume is unchanged. d. its density decreases.

Gas Laws Test Answers

RD Boyd



Gas Laws Test Answers:

The book delves into Gas Laws Test Answers. Gas Laws Test Answers is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Gas Laws Test Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Gas Laws Test Answers
- Chapter 2: Essential Elements of Gas Laws Test Answers
- Chapter 3: Gas Laws Test Answers in Everyday Life
- Chapter 4: Gas Laws Test Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Gas Laws Test Answers. The first chapter will explore what Gas Laws Test Answers is, why Gas Laws Test Answers is vital, and how to effectively learn about Gas Laws Test Answers.
3. In chapter 2, this book will delve into the foundational concepts of Gas Laws Test Answers. This chapter will elucidate the essential principles that must be understood to grasp Gas Laws Test Answers in its entirety.
4. In chapter 3, this book will examine the practical applications of Gas Laws Test Answers in daily life. This chapter will showcase real-world examples of how Gas Laws Test Answers can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Gas Laws Test Answers in specific contexts. The fourth chapter will explore how Gas Laws Test Answers is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Gas Laws Test Answers. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Gas Laws Test Answers.

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