

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Instructions

Standard Conditions

Unit	Standard Pressure
atm	1.013
mmHg	760
gms	1
cm	1.013

22.4 L = 1 mol at STP

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law	Charles's Law
$P_1V_1 = P_2V_2$	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law	Avogadro's Law
$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law	Gramm's Law
$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$

This constant R (used in the ideal gas law) depends on the unit of pressure:

- if kPa, $R = 8.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$
- if atm, $R = 0.0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}}$

Notes:

- Temperature must always be in Kelvin! $K = ^\circ C + 273$
- If you are given two different units of pressure, convert (only)
- Create two columns before solving: initial and final. Write what you're given in the problem. Find the two that use the variable you're given, and solve for the unknown. (for Ideal Gas Law, write what you're given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE! (MAYBE)
- Gramm's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

Charles' Law $V_1/T_1 = V_2/T_2$

Step 1: Compare each of the problems below. How many variables are given in the problem?

1. The temperature under my refrigerator is about 4°C. The gas in the tank has the temperature of 4°C and is under 1 atm. What is the volume of the gas?
2. A gas is at 1 atm and 100°C. What is the volume of the gas?

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