

**12 • The Gas Laws****THE COMBINED GAS LAW**

In practical terms, it is often difficult to hold any of the variables constant. When there is a change in pressure, volume and temperature, the combined gas law is used.

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \text{or} \quad P_1 \times V_1 \times T_2 = P_2 \times V_2 \times T_1$$

$$K = ^\circ C + 273$$

*See  
Scratch  
paper*

Complete the following chart.

	$P_1$	$V_1$	$T_1$	$P_2$	$V_2$	$T_2$
1	1.50 atm	3.00 L	<del>20.0°C</del> 293 K	2.50 atm		<del>30.0°C</del> 303 K
2	720. torr	256. mL	<del>25.0°C</del> 298 K		250. mL	<del>50.0°C</del> 323 K
3	600. mmHg	2.50 L	<del>22.0°C</del> 325 K	760. mmHg	1.80 L	
4		750. mL	<del>0.00°C</del> 273 K	2.00 atm	500. mL	<del>25.0°C</del> 298 K
5	95.0 kPa	4.00 L		101. kPa	6.00 L	<del>498.°C</del> 471. K or
6	650. torr		<del>100.°C</del> 373 K	900. torr	225. mL	<del>150.°C</del> 423 K
7	850. mmHg	1.50 L	<del>15.0°C</del> 288 K		2.50 L	<del>30.0°C</del> 303 K
8	125. kPa	125. mL		100. kPa	100 mL	<del>75.0°C</del> 348 K

# Gas Laws 3 Answer Key

**Neetin Agrawal**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

### Gas Laws 3 Answer Key:

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Heat Neetin Agrawal, 2017-03-08 This book includes theory solved examples and exercise for chapters of Calorimetry Thermal Expansion Kinetic Theory of Gases Heat Transfer Thermodynamics and Elasticity This book will help students preparing for Board exams after class 12th or equivalent This book will be a complete knowledge house for Heat for students preparing for IIT JEE and other similar competitive exams Best of luck to the students using this book

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