

EXTRA PRACTICE PROBLEMS FOR MODULE #6

1. Nitric acid is manufactured with the following reaction:



How many grams of nitrogen dioxide are required to produce 5.89×10^3 kg HNO_3 in excess water?

2. Tungsten metal, W, is produced by the following reaction:



How many grams of tungsten can be made from 1.0×10^6 g of hydrogen in excess WO_3 ?

3. How many grams of WO_3 would be used in the above reaction under the conditions specified in #2?

4. Carbon tetrachloride is made according to the following reaction:



How many grams of CS_2 are needed to make 16 kg of CCl_4 ?

5. To make a phosphorus fertilizer ($\text{CaH}_4\text{P}_2\text{O}_8$), agricultural companies use the following reaction:



If 2.50×10^3 grams of $\text{Ca}_3\text{P}_2\text{O}_8$ are reacted with excess H_2SO_4 and H_2O , how many grams of fertilizer can be made?

6. In the following reaction:



What mass of Mg must you use to make 5.6×10^4 g of Mg(OH)_2 ?

7. Which of the following are empirical formulas?

a. SO_3

b. $\text{K}_2\text{S}_2\text{O}_3$

c. $\text{C}_6\text{H}_{12}\text{O}_6$

d. P_2H_4

8. Decomposition of 100.0 g of a dark green powder yields 39.7 g K, 27.9 g Mn, and an unknown amount of oxygen. What is the empirical formula?

9. The mass of the molecule above is 197.1 amu. What is the molecular formula?

10. A compound has an empirical formula of CH_2O . If its molecular mass is 60.0 amu, what is its molecular formula?

Exploring Creation With Chemistry Practice Problem Answers

Brendan G. Carr



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