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Date: **October 4, 2012**

Student Exploration: Electron Configuration

Activity B: Atomic radii

Get the Gizmo ready:

- On the PERIODIC TABLE tab, select **Na** (sodium).
- Select the ELECTRON CONFIGURATION tab.



Question: How do the radii of atoms change across a period of the periodic table?

1. **Predict:** Positively charged protons in the nucleus of an atom are attracted to negatively charged electrons.

How do you think the **atomic radii** will change as electrons are added to a shell?

It will decrease

2. **Arrange:** Create a proper electron configuration for sodium. After clicking **Check**, note the **Electron configuration** and the **Atomic radius** now listed at right.

Sodium electron configuration: **1s2 2s2 2p6 3s1**

Atomic radius: **190 Pico meters**

3. **Compare:** Click **Next element**, and then add an electron to the magnesium atom. Click **check**, and record the electron configuration and atomic radius below.

Magnesium electron configuration: **1s2 2s2 2p6 3s2**

Atomic radius: **145 Pico meters**

4. **Gather data:** Create electron configurations for the next six elements. Record the electron configuration and atomic radius of each. (Note: The symbol for picometer is *pm*.)

Element	Number of electrons	Electron configuration	Atomic radius (pm)
Aluminum	13	1s2 2s2 2p6 3s2 3p1	118
Silicon	14	1s2 2s2 2p6 3s2 3p2	111
Phosphorus	15	1s2 2s2 2p6 3s2 3p3	98
Sulfur	16	1s2 2s2 2p6 3s2 3p4	88
Chlorine	17	1s2 2s2 2p6 3s2 3p5	79
Argon	18	1s2 2s2 2p6 3s2 3p6	71

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Ying-Ying Zheng



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