

CHEMISTRY I EQUATION SHEET

Solubility Rules

Solubility Rules			Form a Compound which is	
Negative Ion Plus Positive Ion				
Any neg. ion + group 1a ions Soluble				
Any neg. ion + Ammonium ion Soluble				
Nitrate + any positive ion Soluble				
Acetate	+	any positive ion except Ag ⁺ or Hg ²⁺	Soluble	
bromide or iodide + Any other positive ion	+	Ag ⁺ , Pb ²⁺ , Hg ²⁺ or Cu ⁺	Not Soluble	
Sulfate	+	Ca ²⁺ , Sr ²⁺ , Ba ²⁺ , Ra ²⁺ , Ag ⁺ or Pb ²⁺	Not Soluble	
	+	Any other positive ion	Soluble	
Sulfide	+	group 1a ions or ammonium	Soluble	
	+	Be ²⁺ , Mg ²⁺ , Ca ²⁺ , Sr ²⁺ , Ba ²⁺ or Ra ²⁺	Soluble	
	+	Any other positive ion	Not Soluble	
Hydroxide	+	group 1a ions or ammonium	Soluble	
	+	Any other positive ion	Not Soluble	
Carbonate, Chromate, Phosphate, or Sulfite + Any other positive ion	+	group 1a ions or ammonium	Soluble	
	+	Any other positive ion	Not Soluble	

Equations

D = m/v	c = specific heat
Molarity = $\frac{\text{mol}}{\text{V}}$	d = density
$P_1 V_1 = P_2 V_2$	I = initial amount
$\frac{T_1}{T_2}$	F = final amount
Q = mc ΔT	m = mass
nRTPV =	M = molarity
F = I (1/2) *	n = number of moles & number of half lives
n = Total Time	P = pressure
Half Life Time	R = ideal gas constant
	Q = heat energy
	T = temperature
	V = volume
	X = amount of matter left

Constants

	1 atm = 101.3 kPa Pressure
	1 atm = 760 mm Hg
	R = 8.314 L x kPa / mol°C Universal
Constant or 0.0821 L x atm / mol°C	
Avogadro's number	6.02 X
Specific Heats (J / g°C)	
	Water = 4.180
	Aluminum = 0.895
	Copper = 0.387
	Iron = 0.448

Activity Series in Decreasing Activity

Metals
Lithium
Rubidium
Potassium
Calcium
Sodium
Magnesium
Aluminum
Manganese
Zinc
Iron
Nickel
Tin
Lead
Hydrogen
Copper
Silver
Platinum
Gold
Nonmetals
Fluorine
Chlorine
Bromine
Iodine

Polyatomic Ions

Name	Formula
Ammonium	NH ₄ ⁺
Acetate	C ₂ H ₃ O ₂ ⁻
Carbonate	CO ₃ ⁻²
Bicarbonate or Hydrogen carbonate	HCO ₃ ⁻
Hypochlorite	ClO ⁻
Chlorite	ClO ₂ ⁻
Chlorate	ClO ₃ ⁻
Perchlorate	ClO ₄ ⁻
Hydroxide	OH ⁻
Cyanide	CN ⁻
Nitrate	NO ₃ ⁻
Nitrite	NO ₂ ⁻
Chromate	CrO ₄ ⁻²
Oxalate	C ₂ O ₄ ⁻²
Sulfate	SO ₄ ⁻²
Sulfite	SO ₃ ⁻²
Hydrogen sulfate	HSO ₄ ⁻
Phosphate	PO ₄ ⁻³
Phosphite	PO ₃ ⁻³
Hydrogen phosphate	HPO ₄ ⁻²

Metric Conversions

Unit	Abbr.	Meaning
Mega-	M	10 ⁶
kilo-	k	10 ³
hecto-	h	10 ²
deca-	da	10 ¹
Liter	L	1
gram	g	1
meter	m	1
deci-	d	10 ⁻¹
centi-	c	10 ⁻²
milli-	m	10 ⁻³
micro-	μ	10 ⁻⁶
nano-	n	10 ⁻⁹

Calorie Conversions

Water = 1 cal/ g°C
1000 cal = 1 Cal

High School Chemistry Formulas Cheat Sheet

Annelies Wilder-Smith



High School Chemistry Formulas Cheat Sheet:

Reviewing **High School Chemistry Formulas Cheat Sheet**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**High School Chemistry Formulas Cheat Sheet**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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