



**Western  
Science**

## Chemistry 1302B Final Exam \*Answer Key\*

Wednesday, April 19, 2023  
2:00 – 5:00 pm

1. The density of the  $\text{N}_2(\text{g})$  in a rigid tank at 250 kPa was  $1.54 \text{ g L}^{-1}$ . What was the temperature of the gas in this container?

547 K

2. Which of the following statements is true regarding the two samples below?

- Sample I: 3.0 moles of Ar at 20 °C in a container of volume V
- Sample II: 3.0 moles of Ar at 100 °C in a container of volume V

The pressure of Sample II is greater than that of Sample I.

3. Before contact with Europeans, Indigenous Peoples in this region produced maple syrup by evaporating the water from maple sap held in wooden containers. To heat the sap, they would drop in some stones that had been heated in a fire. If a 500 g stone that had been heated to 200 °C was added to a container holding 2.00 kg of sap at 5 °C, what temperature would the liquid have achieved when stone and sap reached thermal equilibrium? (Helpful data: The specific heat capacity of the stone was  $0.790 \text{ J g}^{-1} \text{ K}^{-1}$  and that of the sap at this stage was the same as that of liquid water. Ignore heat losses to the container and surroundings.)

53.8 °C

4. Given:  $\Delta H_f^\circ(\text{CO}_2\text{g}) = -130 \text{ kJ mol}^{-1}$ , the  $\Delta H^\circ$  of the reaction  $\text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{CO}_2(\text{g})$  is:

- I.  $\Delta H_f^\circ$  of  $\text{CO}_2(\text{g})$
- II. The heat of combustion of  $\text{CO}(\text{g})$
- III.  $-504 \text{ kJ}$

II only

5. Given:  $\Delta H_{\text{f, H}_2\text{O}} = -422.6 \text{ kJ mol}^{-1}$        $\Delta H_{\text{f, CO}} = 535.6 \text{ kJ mol}^{-1}$

Estimate  $\Delta H$  for the reaction:  $3\text{H}_2(\text{g}) + \text{CO}_2(\text{g}) \rightarrow 3\text{H}_2\text{O}(\text{g}) + \text{C}(\text{s})$

-246 kJ

6. How many of the processes below have a positive  $\Delta S^\circ$ ?

- $\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightarrow 2 \text{NH}_3(\text{g})$
- $2 \text{NaCl}(\text{s}) + \text{Cl}_2(\text{g}) \rightarrow 2 \text{NaCl}(\text{aq})$
- Heating 10 g of water from 0 °C to 10 °C
- Condensing 10 g of water vapour into liquid at 100 °C

# Gases Section 1 Answers

**Patrick Vollmar**

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