

1. A gas has a pressure of 0.370 atm at 55.0 °C. What is the pressure at standard temperature?

$$P_2 = P_1 \frac{T_2}{T_1} = \frac{(0.370 \text{ atm})(273 \text{ K})}{328 \text{ K}} = 0.313 \text{ atm}$$

2. If 2.00 mole of gas occupies 4.50L at STP. How much of the same gas will occupy 3.00L at STP?

$$n_2 = n_1 \frac{V_2}{V_1} = \frac{(2.00 \text{ mole})(3.00 \text{ L})}{4.50 \text{ L}} = 1.33 \text{ mole}$$

3. I've got a car with an internal volume of 12,000 L. If I drive my car into the river and it implodes, what will be the volume of the gas when the pressure goes from 1.0 atm to 1.4 atm?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(1.0 \text{ atm})(12,000 \text{ L})}{1.4 \text{ atm}} = 8571 \text{ L}$$

4. What's the partial pressure of carbon dioxide in a container that holds 5.00 moles of carbon dioxide, 3.00 moles of nitrogen, and 1.00 mole of hydrogen and has a total pressure of 1.05 atm?

$$P_{\text{CO}_2} = P_T \times \frac{\text{moles CO}_2}{\text{Total moles}} = (1.05 \text{ atm}) \times \frac{(5.00 \text{ moles})}{(9.00 \text{ moles})} = 0.583 \text{ atm}$$

5. I have 130 liters of gas in a piston at a temperature of 250. °C. If I cool the gas until the volume decreases to 85 liters, what will temperature of the gas be?

$$T_2 = \frac{V_2 T_1}{V_1} = \frac{(85 \text{ L})(250 \text{ K})}{130 \text{ L}} = 162 \text{ K} = -111 \text{ °C}$$

6. If a gas in a closed container is pressurized from 15.0 atmospheres to 16.0 atmospheres and its original temperature was 25.0 °C, what would the final temperature of the gas be?

$$T_2 = \frac{P_2 T_1}{P_1} = \frac{(16.0 \text{ atm})(298 \text{ K})}{15.0 \text{ atm}} = 318 \text{ K}$$

7. To compress nitrogen at 1 atm from 750 mL to 500 mL, what must the new pressure be if the temperature is kept constant?

$$P_2 = \frac{V_1 P_1}{V_2} = \frac{(750 \text{ mL})(1 \text{ atm})}{500 \text{ mL}} = 1.5 \text{ atm}$$

8. Calcium carbonate decomposes at 1200. °C to form carbon dioxide and calcium oxide. If 25 liters of carbon dioxide are collected at 1200. °C, what will the volume of this gas be after it cools to 25 °C?

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{1473 \text{ K}} = 5.1 \text{ L}$$

9. 5.00 L of a gas is known to contain 0.065 moles. If this amount of gas is increased to 1.80 moles, what new volume will result (at an unchanged temperature and pressure)?

$$V_2 = \frac{V_1 n_2}{n_1} = \frac{(5.00 \text{ L})(1.80 \text{ moles})}{0.065 \text{ moles}} = 138 \text{ L}$$

Gas Law Worksheet 1 Answers

E Durkheim



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A Guide to Teaching in the Active Learning Classroom Paul Baeppler, J. D. Walker, D. Christopher Brooks, Kem Saichaie, Christina I. Petersen, 2023-07-03 While Active Learning Classrooms or ALCs offer rich new environments for learning they present many new challenges to faculty because among other things they eliminate the room s central focal point and disrupt the conventional seating plan to which faculty and students have become accustomed The importance of learning how to use these classrooms well and to capitalize on their special features is paramount The potential they represent can be realized only when they facilitate improved learning outcomes and engage students in the learning process in a manner different from traditional classrooms and lecture halls This book provides an introduction to ALCs briefly covering their history and then synthesizing the research on these spaces to provide faculty with empirically based practical guidance on how to use these unfamiliar spaces effectively Among the questions this book addresses are How can instructors mitigate the apparent lack of a central focal point in the space What types of learning activities work well in the ALCs and take advantage of the affordances of the room How can teachers address familiar classroom management challenges in these unfamiliar spaces If assessment and rapid feedback are critical in active learning how do they work in a room filled with circular tables and no central focus point How do instructors balance group learning

with the needs of the larger class How can students be held accountable when many will necessarily have their backs facing the instructor How can instructors evaluate the effectiveness of their teaching in these spaces This book is intended for faculty preparing to teach in or already working in this new classroom environment for administrators planning to create ALCs or experimenting with provisionally designed rooms and for faculty developers helping teachers transition to using these new spaces

Simplified ICSE Chemistry Viraf J. Dalal, **Concepts of Earth Science & Chemistry Parent Lesson Plan** John Hudson Tiner, 2013-08-26 Concepts of Earth and Chemistry Course Description This is the suggested course sequence that allows one core area of science to be studied per semester You can change the sequence of the semesters per the needs or interests of your student materials for each semester are independent of one another to allow flexibility Semester 1 Earth Blending a creationism perspective of history with definitions of terms and identification of famous explorers scientists etc this book gives students an excellent initial knowledge of people and places encouraging them to continue their studies in depth Semester 2 Chemistry Chemistry is an amazing branch of science that affects us every day yet few people realize it or even give it much thought Without chemistry there would be nothing made of plastic there would be no rubber tires no tin cans no televisions no microwave ovens or something as simple as wax paper This book presents an exciting and intriguing tour through the realm of chemistry as each chapter unfolds with facts and stories about the discoveries of discoverers Find out why pure gold is not used for jewelry or coins Join Humphry Davy as he made many chemical discoveries and learn how they shortened his life See how people in the 1870s could jump over the top of the Washington Monument Exploring the World of Chemistry brings science to life and is a wonderful learning tool with many illustrations and biographical information

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