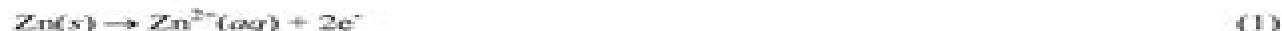


Experiment 9

Electrochemistry I – Galvanic Cell

Introduction: Chemical reactions involving the transfer of electrons from one reactant to another are called *oxidation-reduction* reactions or *redox* reactions. In a redox reaction, two *half-reactions* occur; one reactant gives up electrons (undergoes oxidation) and another reactant gains electrons (undergoes reduction). A piece of zinc going into a solution as zinc ions, with each Zn atom giving up 2 electrons, is an example of an oxidation half-reaction.



The oxidation number of $\text{Zn}(s)$ is 0 and the oxidation number of the Zn^{2+} is +2. Therefore, in this half-reaction, **the oxidation number increases**, which is another way of defining an oxidation. In contrast, the reverse reaction, in which Zn^{2+} ions gain 2 electrons to become Zn atoms, is an example of reduction.



In a reduction there is a **decrease (or reduction) in oxidation number**. Chemical equation representing half-reactions must be both mass and charge balanced. In the half-reactions above, there is one zinc on both sides of the equation. The charge is balanced because the 2+ charge on the zinc ion is balanced by two electrons, $2e^{-}$, giving zero net charge on both sides.

Another example of reduction is the formation of solid copper from copper ions in solution.



In this half-reaction the oxidation number of the aqueous copper is +2, which decreases to 0 for the solid copper, and again charge and mass are balanced. However, no half-reaction can occur by itself. A redox reaction results when an oxidation and a reduction half-reaction are combined to complete a transfer of electrons as in the following example:



The electrons are not shown because they are neither reactants nor products but have simply been transferred from one species to another (from Zn to Cu^{2+} in this case). In this redox reaction, the $\text{Zn}(s)$ is referred to as the *reducing agent* because it causes the Cu^{2+} to be reduced to Cu. The Cu^{2+} is called the *oxidizing agent* because it causes the $\text{Zn}(s)$ to be oxidized to Zn^{2+} .

Any half-reaction can be expressed as a reduction as illustrated in the case where equation (1) can be reversed to equation (2). A measure of the tendency for a reduction to occur is its *reduction potential*, E , measured in units of volts. At standard conditions, 25 °C and concentrations of 1.0 M for the aqueous ions, the measured voltage of the reduction half-reaction is defined as the *standard reduction potential*, E° . Standard reduction potentials

Galvanic Cell Experiment Report

J Elliott



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