

Roll No. . . . .

Oct., 2022/10121

**B.Sc. Nursing I<sup>st</sup> Semester-2022**  
**Applied Anatomy & Applied Physiology**

Time: Three hours

M. Marks: 75

- Note: 1. Attempt all of the following questions and draw suitable diagrams where required.  
 2. Attempt Part-I & Part-II in separate answer books.  
 3. Use only blue/black ink pen for attempting answers. Use of pencil is prohibited.

**Part-I (Applied Anatomy)**

**Q.1 Multiple choice questions:**

6x1=6

- Most common vein used for intravenous injections:  
 (a) Cephalic vein (b) Median cubital vein  
 (c) Basilic vein (d) Radial artery
- Which is not a layer of skin:  
 (a) Epidermis (b) Dermis  
 (c) Hypodermis (d) Stratum corneum
- Membranous ossification shown by:  
 (a) Humerus (b) Patella  
 (c) Femur (d) Frontal
- Which is not an action of deltoid:  
 (a) Flexion (b) Extension  
 (c) Adduction (d) Abduction
- Inner lining of uterine cavity:  
 (a) Endothelium (b) Endometrium  
 (c) Endocardium (d) Myometrium
- Islets of langerhans seen in which organ:  
 (a) Spleen (b) Small intestine  
 (c) Pancreas (d) Stomach

**Q.2 Describe the organ heart under the following headings:**

3+2+5=10

- Chambers.
- Valves.
- Blood supply.

**Q.3 Write short note on the following:**

3x5=15

- Types of cartilage.
- CSF circulation.
- Internal structure of kidney.

**Q.4 Very short answer type questions:**

3x2=6

- Enumerate the muscles of posterior compartment of leg.
- Enumerate the branches of external carotid artery.
- Draw a well labeled diagram of multipolar neuron.

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**O García**



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jan 3 2023 a net ionic equation is a chemical equation that shows only the species participating in a chemical reaction and omits the spectator ions spectator ions are ions that

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is  $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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jun 23 2022 what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic

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aug 10 2023 write the full ionic and net ionic equations for this reaction solution the elemental metals magnesium on the reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative

descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds  $\text{Cl}^-\text{aq}$

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver  $\text{I}^+$  ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4  $13\text{Ag}^+\text{aq} + 2\text{F}^-\text{aq} + 2\text{NH}_4^+\text{aq} + \text{Cr}_2\text{O}_7^{2-}\text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+\text{aq} + 2\text{F}^-\text{aq}$  because two  $\text{NH}_4^+\text{aq}$

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