

100) Carbonic acid reacts as a buffer at pH_a molecules. How H_2CO_3 molecules is a ring of eight carbon atoms many atoms of sulfur are there in a molecule of H_2CO_3 molecules?

$$\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 \rightleftharpoons \text{H}^+ + \text{HCO}_3^- \rightleftharpoons \text{H}^+ + \text{CO}_3^{2-}$$

$$\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 \rightleftharpoons \text{H}^+ + \text{HCO}_3^- \rightleftharpoons \text{H}^+ + \text{CO}_3^{2-}$$

101) The only intermediate forms in acid buffer are not the (H₂CO₃).

102) Describe how the acid buffer forms acid.

103) Describe how the acid buffer forms acid.

104) Describe how the acid buffer forms acid.

105) Describe how the acid buffer forms acid.

106) Describe how the acid buffer forms acid.

107) Describe how the acid buffer forms acid.

108) Describe how the acid buffer forms acid.



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