

AP QUESTION - 1995

(a) mole fraction CO
assume 1.0 L



$$Y_{\text{CO}} = \frac{.55 \text{ mol CO}}{1.6 \text{ mol total}} = .34375 = \boxed{.34}$$

$$(b) K_c = \frac{[\text{H}_2\text{O}][\text{CO}]}{[\text{H}_2][\text{CO}_2]} = \frac{(.55 \text{ M})(.55 \text{ M})}{(.20 \text{ M})(.30 \text{ M})} = \boxed{5.04} \quad (\text{no unit})$$

$$(c) K_p = K_c (RT)^{\Delta n} \quad \Delta n = 0 \quad \therefore K_p = K_c = \boxed{5.04}$$

(d)

$$\text{H}_2 + \text{CO}_2 \rightleftharpoons \text{H}_2\text{O} + \text{CO}$$

i	.20 M	.30 M	.55 M	.55 M
Δ	+ .165 M	+ .165 M	- .165 M	- .165 M
eq	.365	.465	.385	.385

$$30\% \text{ of } .55 = .165$$

$$K_c = \frac{[\text{H}_2\text{O}][\text{CO}]}{[\text{H}_2][\text{CO}_2]} = \frac{(.385)(.385)}{(.365)(.465)} = \boxed{.873}$$

(e)

i	.50 / $\sqrt{2}$.50 / $\sqrt{2}$	0 M	0 M
Δ	-x	-x	+x	+x
eq	.167 - x	.167 - x	x	x

$$\frac{x^2}{(.167 - x)^2} = 5.04$$

$$\frac{x}{.167 - x} = \sqrt{5.04}$$

$$\frac{x}{.167 - x} = 2.24$$

$$\boxed{x = .116 \text{ M CO}}$$

TAKE
R BOTH
SIDES

Exploring Gas Laws Simulation Lab Answers

Mark Steyvers



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Table of Contents Exploring Gas Laws Simulation Lab Answers

1. Understanding the eBook Exploring Gas Laws Simulation Lab Answers
 - The Rise of Digital Reading Exploring Gas Laws Simulation Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Exploring Gas Laws Simulation Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Exploring Gas Laws Simulation Lab Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Exploring Gas Laws Simulation Lab Answers
 - Personalized Recommendations
 - Exploring Gas Laws Simulation Lab Answers User Reviews and Ratings
 - Exploring Gas Laws Simulation Lab Answers and Bestseller Lists
5. Accessing Exploring Gas Laws Simulation Lab Answers Free and Paid eBooks
 - Exploring Gas Laws Simulation Lab Answers Public Domain eBooks
 - Exploring Gas Laws Simulation Lab Answers eBook Subscription Services
 - Exploring Gas Laws Simulation Lab Answers Budget-Friendly Options
6. Navigating Exploring Gas Laws Simulation Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Exploring Gas Laws Simulation Lab Answers Compatibility with Devices
 - Exploring Gas Laws Simulation Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Exploring Gas Laws Simulation Lab Answers
 - Highlighting and Note-Taking Exploring Gas Laws Simulation Lab Answers
 - Interactive Elements Exploring Gas Laws Simulation Lab Answers
8. Staying Engaged with Exploring Gas Laws Simulation Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Exploring Gas Laws Simulation Lab Answers
9. Balancing eBooks and Physical Books Exploring Gas Laws Simulation Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Exploring Gas Laws Simulation Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Exploring Gas Laws Simulation Lab Answers
 - Setting Reading Goals Exploring Gas Laws Simulation Lab Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Exploring Gas Laws Simulation Lab Answers
 - Fact-Checking eBook Content of Exploring Gas Laws Simulation Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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