



Figure 2

4. Collect pressure data at three different temperatures (cold tap water, ice water, and warm tap water). Set up water baths in a 600 mL beaker. Take a picture (or print a copy) of the graph and data.

Initial by a TA or an instructor (for data and graph): me

### Post-lab Questions:

1. Attach the data and graphs for Part I and Part II.
2. For each of the two parts of the experiment, write an equation using the two variables and a proportionality constant,  $k$  (e.g., for Part I,  $P = k \times V$  if direct, or  $P = k/V$  if inverse).
3. Calculate the constant,  $k$ , for each of the two gas laws that you tested. This value can be an average for each of the data pairs in each part of the experiment.
4. For Part I, what would the pressure be if the volume were 0.3 mL? (Use your average  $k$  for Part I from question #3.)
5. For Part II, what would be the pressure if the temperature were 100 °C? (Use your average  $k$  for Part II from question #3.)

# Exploring The Properties Of Gases Lab Answers

**M Mosston**



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