



Group members:

101

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Practical Result: K_c for the reaction

✓

✓

✓

✓

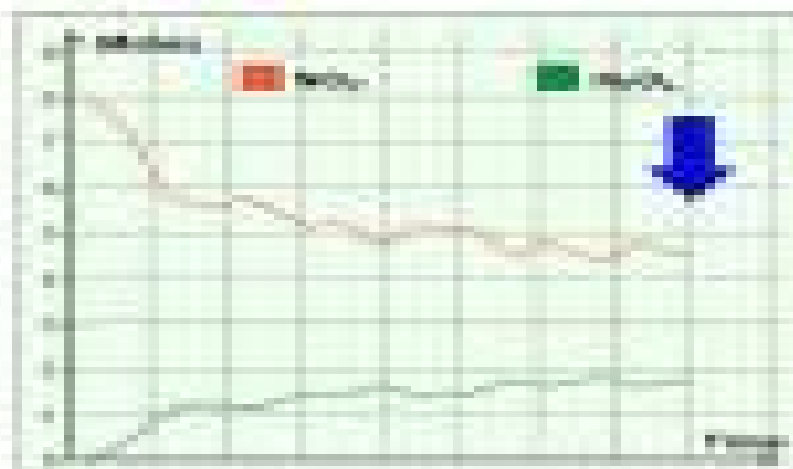
✓

10/10

Practical Skills: 10/10

Question 1: Answer the following

1. A closed flask containing hydrogen gas at a pressure of 1 atm is heated to 100°C. The flask is then cooled to 0°C. Calculate the change in the equilibrium constant K_c for the reaction $2H_2(g) \rightleftharpoons 2H(g)$ at 100°C and 0°C.



- K_c is independent of temperature and is constant for all reactions.
- The forward reaction is exothermic and the reverse reaction is endothermic.
- The rate of the forward reaction, k_f , is greater than the rate of the reverse reaction, k_r , at equilibrium.
- The activation energy for the forward reaction is less than that for the reverse reaction.

Answer: (a) The value of the equilibrium constant K_c is independent of temperature and is constant for all reactions.

(b) The forward reaction is exothermic and the reverse reaction is endothermic. The value of the equilibrium constant K_c is greater than 1 at 100°C and less than 1 at 0°C.

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