

Name: _____

Date: _____

Student Exploration: Covalent Bonds

Activity B: Building molecules	Get the Gizmo ready: • Click Reset. • Turn off Show Lewis diagram. • Select Oxygen.	
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Question: How do atoms share more than one pair of electrons?

- Observe:** Like fluorine and most other elements, oxygen atoms are most stable with a full complement of eight valence electrons.
 - How many valence electrons does each oxygen atom have now? 6
 - How many more electrons does each oxygen atom need to be stable? 2
- Form a bond:** Drag electrons back and forth until the molecule of oxygen (O_2) is stable. Click **Check** to confirm your molecule is stable. Take a snapshot and paste the image into your document (don't forget to label it " O_2 ").

How many *pairs* of shared electrons are there in a stable molecule of oxygen? 1

- Draw a diagram:** Draw a Lewis diagram of the oxygen molecule in the space below at left. To check your work, turn on **Show Lewis diagram**. Draw the correct diagram on the right.

Practice diagram: $\cdot\ddot{O} \equiv \ddot{O}\cdot$

Actual: $\cdot\ddot{O} \equiv \ddot{O}\cdot$

- Practice:** Create covalent bonds and stable molecules for the remaining substances. Take a snapshot of each completed molecule and add it to your document. Draw Lewis diagrams for each one. (As above, draw the diagram on your own before checking your work.)

Nitrogen $\cdot N \equiv N \cdot$

Ammonia



Chlorine $\cdot\ddot{Cl} - \ddot{Cl}\cdot$

Methane



Water $H - \underset{\cdot}{\underset{\cdot}{O}} - H$

Carbon dioxide $\cdot\ddot{O} \equiv C \equiv \ddot{O}\cdot$

Formaldehyde



(Activity B continued on next page)

Explore Learning Covalent Bonds Answer Key

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