

- (b) (i) Based on Maxwell's relations, prove that for any fluid,
- $$dS = C_v \frac{dT}{T} + \left( \frac{\partial p}{\partial T} \right)_v dV. \quad (8)$$

- (ii) A mixture of 3 moles of helium, 4 moles of neon and 5 moles of argon is at 1 bar and 300°K. Calculate

- (1) volume,
- (2) mole fraction and partial pressure of gases and
- (3) change of entropy due to mixing. (8)

15. (a) The moist air is at 45°C Dry Bulb Temperature and 30°C Wet Bulb Temperature. Calculate

- (i) Vapour pressure
- (ii) Dew point temperature
- (iii) Specific enthalpy
- (iv) Relative humidity
- (v) Degree of saturation
- (vi) Vapour density
- (vii) Enthalpy of mixture. (16)

Or

- (b) (i) Explain the process of cooling with dehumidification of air, with the help of sketches of schematic of air duct with the elements involved as well as the representation on psychrometric chart. (6)

- (ii) An air-water vapour mixture at 20°C and 50% relative humidity at a pressure of 1.013 bar is heated at constant pressure to a temperature of 35°C. Calculate

- (1) The initial and final specific humidity
- (2) Final relative humidity
- (3) Dew point temperature
- (4) Heat transferred per kg of dry air. (10)

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## **Table of Contents Engineering Thermodynamics Question Paper 2012**

1. Understanding the eBook Engineering Thermodynamics Question Paper 2012
  - The Rise of Digital Reading Engineering Thermodynamics Question Paper 2012
  - Advantages of eBooks Over Traditional Books
2. Identifying Engineering Thermodynamics Question Paper 2012
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Engineering Thermodynamics Question Paper 2012
  - User-Friendly Interface
4. Exploring eBook Recommendations from Engineering Thermodynamics Question Paper 2012
  - Personalized Recommendations
  - Engineering Thermodynamics Question Paper 2012 User Reviews and Ratings
  - Engineering Thermodynamics Question Paper 2012 and Bestseller Lists
5. Accessing Engineering Thermodynamics Question Paper 2012 Free and Paid eBooks
  - Engineering Thermodynamics Question Paper 2012 Public Domain eBooks
  - Engineering Thermodynamics Question Paper 2012 eBook Subscription Services
  - Engineering Thermodynamics Question Paper 2012 Budget-Friendly Options

6. Navigating Engineering Thermodynamics Question Paper 2012 eBook Formats
  - ePub, PDF, MOBI, and More
  - Engineering Thermodynamics Question Paper 2012 Compatibility with Devices
  - Engineering Thermodynamics Question Paper 2012 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Engineering Thermodynamics Question Paper 2012
  - Highlighting and Note-Taking Engineering Thermodynamics Question Paper 2012
  - Interactive Elements Engineering Thermodynamics Question Paper 2012
8. Staying Engaged with Engineering Thermodynamics Question Paper 2012
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Engineering Thermodynamics Question Paper 2012
9. Balancing eBooks and Physical Books Engineering Thermodynamics Question Paper 2012
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Engineering Thermodynamics Question Paper 2012
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Engineering Thermodynamics Question Paper 2012
  - Setting Reading Goals Engineering Thermodynamics Question Paper 2012
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Engineering Thermodynamics Question Paper 2012
  - Fact-Checking eBook Content of Engineering Thermodynamics Question Paper 2012
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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