

Accelerated Chemistry – Test Study Guide Liquids and Solids: Chapter 10

- Use Particle Theory to describe and contrast the 3 states of matter in-terms of particle arrangement, particle speed, diffusion rate, volume and shape
- List, define and compare the physical properties of liquids and solids
- Describe the affect of intermolecular forces (IMF) of attraction on the phases of matter and their properties
- List and explain the reason for the unique and unusual properties of water
- What is surface tension? What affects the surface tension of a liquid?
- Define and give examples of crystalline solids and amorphous solids
- List and define the changes of states in matter
- What is the “equilibrium”? How is it represented with a symbol?
- What is equilibrium vapor pressure of a liquid?
- What factors affect the vapor pressure of a liquid and why?
- Read and interpret vapor pressure curves for two different liquids
- What is a volatile liquid? Give an example?
- Explain the difference between boiling and evaporation
- What is meant by the “*normal*” boiling point of a liquid?
- How is boiling point affected by atmospheric pressure?
- Read and label a phase diagram for a pure substance
- Predict the changes in state on a phase diagram due changes in T and P
- What is the triple point and critical temperature on a phase diagram?
- Calculate the heat required to cause a single phase change in matter
- Calculate the total heat required to convert ice to steam
- Know how to sketch, label and read a heating curve

Equations to know how to use:

Heat for a temperature change: $H = mc\Delta T$ m = mass, c = specific heat, $\Delta T = T_f - T_i$

Heat to melt a solid: $H_m = n \cdot \Delta H_{fus}$ n = # moles, ΔH_{fus} = molar heat of fusion

Heat to vaporize a liquid: $H_v = n \cdot \Delta H_{vap}$ n = # moles, ΔH_{vap} = molar heat of Vaporization

Liquids And Solids Chemistry Study Guide Answers

Yi-Tong Ma



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