

Question 5

1. A chemical reaction involving protons and hydroxide ions is shown below. Calculate the equilibrium constant K_a for the reaction at 25°C. The concentration of the products is 0.1 M and the concentration of the reactants is 0.1 M.

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

2. Calculate the pH of a 0.1 M solution of $H_2PO_4^-$.
Give your answer to two decimal places.

$$H^+ = \sqrt{(4.0 \times 10^{-8})(0.1)} = 0.0002$$

$$-\log(0.0002) = 3.70$$

$$pH = 3.70$$

3. The chemical reaction below is a buffer reaction. Calculate the pH of a buffer solution, with 0.1 M of $H_2PO_4^-$ and 0.1 M of HPO_4^{2-} .

4. Calculate the pH of the buffer solution.

Give your answer to two decimal places.

$$0.1 \times \frac{10^{-7}}{10^{-7}} = 0.01$$

$$0.01 \times \frac{10^{-7}}{10^{-7}} = 0.01$$

$$(4.6 \times 10^{-8}) \times \frac{0.01}{0.01} = 0.0000046$$

$$-\log(0.0000046) = 5.34$$

$$5.34$$

5. A chemical reaction involving protons and hydroxide ions is shown below. Calculate the equilibrium constant K_a for the reaction at 25°C. The concentration of the products is 0.1 M and the concentration of the reactants is 0.1 M.

6. Calculate the pH of a 0.1 M solution of $H_2PO_4^-$.
Give your answer to two decimal places.

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$$-\log(0.0000046) = 5.34$$

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